

Acids and bases review sheet Answer key

1. List the differences between an acid and a base. Your description should include such things as whether both are electrolytes, which contains a hydrogen ion and which includes a hydroxide ion, and you should also include two examples of an acid and two examples of a base that are present in your household.

| Acid | base |
|------------------------------|-------------------------|
| Conducts electricity | conducts electricity |
| Contains hydrogen ion | contains hydroxide ion |
| Proton donor | proton acceptor |
| Turns litmus paper red | turns litmus paper blue |
| React with most metals | releases hydroxide ion |
| Releases hydronium | Ex: Windex |
| Ex: lemon | bleach |
| Vinegar | soap |
| Battery acid (sulfuric acid) | |

2. Listed below are some of the properties of acids and bases. Fill in the blanks with the appropriate word, acids or bases:

- a) A produce hydrogen ions (H^+) in solution
- b) B have a bitter taste
- c) B produce hydroxide ions (OH^-) ions
- d) A react with bases to form salts
- e) A have a sour taste
- f) B have a slippery, soapy feel
- g) B react with acids to form salts
- h) A react with many metals

3. Write an equation for the dissociation of the following acids in water: follow example

- a) $HClO_4 \rightarrow H^+ + ClO_4^-$
- b) $H_2SO_4 \rightarrow 2H^+ + SO_4^{2-}$
- c) $HC_2H_5O_2 \rightarrow H^+ + C_2H_5O_2^-$
- d) $H_2S \rightarrow 2H^+ + S^{2-}$
- e) $HCl \rightarrow H^+ + Cl^-$
- f) $HNO_3 \rightarrow H^+ + NO_3^-$

4. Write an equation for the dissociation (ionization) of the following bases in water: (follow example)

- a) $Mg(OH)_2 \rightarrow Mg^{2+} + 2OH^-$
- b) $NaOH \rightarrow Na^+ + OH^-$
- c) $Ba(OH)_2 \rightarrow Ba^{2+} + 2(OH^-)$
- d) $KOH \rightarrow K^+ + OH^-$
- e) $Mg(OH)_2 \rightarrow Mg^{2+} + 2(OH^-)$
- f) $LiOH \rightarrow Li^+ + OH^-$

5) A. What type of reaction is a neutralization reaction? Double replacement

B. Write the net equation for neutralization:



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J Dewey

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