

Comprehensive Study Guide on Acid and Base Equilibria

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Introduction to Acids and Bases

- **Bronsted-Lowry Definition:**
 - **Acid:** H^+ donor
 - **Base:** H^+ acceptor
- **Strong Acid:** Completely ionized in aqueous solution.
- **Weak Acid:** Only partially ionizes.

Key Concepts

- The equilibrium constant (K) does not include the solvent.

Acid Strength and Molecular Structure

- **K_a (Acid Ionization Constant):**
 - $K_a > 1$ indicates a strong acid (complete ionization).
- **Strong Acids Examples:**

Acid Base Equilibria Study Guide Answers

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