

Acid Dissociation Constant (Ka)

Acid Name	Acid Formula	Acid Dissociation Constant (Ka)	pKa
Perchloric acid	HClO_4	1.6×10^{11}	-15.2
Hydroiodic acid	HI	3.2×10^2	-9.5
Hydrobromic acid	HBr	1×10^2	-9
Hydrochloric acid	Cl	1.0×10^1	-6.11
Sulfuric acid	H_2SO_4	1.0×10^2	-3
Hydronium	H_3O^+	1	0
Citric acid	$\text{C}_6\text{H}_5\text{O}_7$	7.4×10^{-5}	3.13
Hydrofluoric acid	HF	6.6×10^{-4}	3.18
Formic acid	HCOOH	1.8×10^{-4}	3.74
Benzoic acid	$\text{C}_6\text{H}_5\text{COOH}$	6.3×10^{-5}	4.20
Acetic acid	CH_3COOH	1.8×10^{-5}	4.74
Carbonic acid	H_2CO_3	4.3×10^{-7}	6.37
Hydrogen sulfide	H_2S	9.5×10^{-8}	7.02
Hypochlorous acid	HOCl	3.5×10^{-8}	7.46
Phenol	$\text{C}_6\text{H}_5\text{OH}$	1.0×10^{-10}	10

Acids dissociate more \Rightarrow Stronger acids



$$K_a = \frac{[\text{H}_3\text{O}^+] [\text{A}^-]}{[\text{HA}]}$$

$$pK_a = - \log_{10}(K_a)$$

$[\text{HA}]$: Undissociated acid concentration

$[\text{H}_3\text{O}^+]$: Hydronium ion concentration

$[\text{A}^-]$: Conjugate base concentration

Acids dissociate less \Rightarrow Weaker acids

Acid 4 198

Julia Schneider

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