

Acid Name	Acid Formula	Acid Dissociation Constant (Ka)	pKa
Perchloric acid	HClO ₄	1.6 x 10 ¹³	-15.2
Hydroiodic acid	HI	3.2 x 10 ⁹	-9.5
Hydrobromic acid	HBr	1 x 10 ⁹	-9
Hydrochloric acid	Cl	1.0 x 10 ⁶	-6.11
Sulfuric acid	H ₂ SO ₄	1.0 x 10 ³	-3
Hydronium	H₃O⁺	1	0
Citric acid	C ₆ H ₈ O ₇	7.4 x 10 ⁻⁴	3.13
Hydrofluoric acid	HF	6.6 x 10 ⁻⁴	3.18
Formic acid	HCOOH	1.8 x 10 ⁻⁴	3.74
Benzoic acid	C ₆ H ₅ COOH	6.3 x 10 ⁻⁵	4.20
Acetic acid	CH ₃ COOH	1.8 x 10 ⁻⁵	4.74
Carbonic acid	H ₂ CO ₃	4.3 x 10 ⁻⁷	6.37
Hydrogen sulfide	H ₂ S	9.5 x 10 ⁻⁸	7.02
Hypochlorous acid	HOCl	3.5 x 10 ⁻⁸	7.46
Phenol	C ₆ H ₅ OH	1.0 x 10 ⁻¹⁰	10



Acids dissociate more => Stronger acids



$$K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$$

$$\text{p}K_a = -\log_{10}(K_a)$$

[HA] : Undissociated acid concentration

[H₃O⁺] : Hydronium ion concentration

[A⁻] : Conjugate base concentration

Acids dissociate less => Weaker acids

Acid 4 198

Julia Schneider



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The Chemical Trade Journal and Chemical Engineer ,1908 **Watts' Dictionary of Chemistry** Henry Watts, Henry Forster Morley, Matthew Moncrieff Pattison Muir,1889 **Drug Metabolism Reviews** ,1985 **Proceedings of the American Pharmaceutical Association** ,1930 *Index to Gmelin's Handbook of Chemistry* ,1872 *The Coal-tar-colours* ,1896 **Hand-book of Chemistry** Cavendish Society (London),1872 **Works of the Cavendish Society: Gmelin, Leopold. Hand-book of chemistry. 18 v. & index. 1848-72** ,1872

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