

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from pH and hydroxide ion concentrations from pOH.
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|-----------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the hydrogen ion concentration or autoionization equilibrium of the solution described. Then calculate the pH or pOH. Round significant figures according to the data. Autoionization is optional.

1. What substance is _____, contains hydrogen ions, and autoionizes? _____ is the autoionization of this acidic substance. Write the equation for _____.
2. Hydrochloric acid has a pH of _____, or a _____ molar _____ concentration. Calculate the _____ of this acid. The _____ is _____.
3. _____, contains the same ions as _____ and _____.
4. _____, contains the same ions as _____ and _____.
5. _____, contains the same ions as _____ and _____.
6. _____, contains the same ions as _____ and _____.
7. _____, contains the same ions as _____ and _____.

8. _____
9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____
17. _____
18. _____

19 2 Hydrogen Ions And Acidity Section Review Answers

**Kevin T. Patton, Gary A.
Thibodeau, Andrew Hutton**



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