

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

- **Crystalline Solids** have a regular, repeating arrangement of particles. They have sharp melting points and distinct X-ray diffraction patterns.
- **Amorphous Solids** lack a long-range order. They soften over a range of temperatures and do not have sharp X-ray diffraction patterns.
- **Intermolecular Forces** are crucial in determining the properties of solids. Stronger forces lead to higher melting points and greater hardness.
- **Hydrogen Bonding** is a special type of dipole-dipole interaction that significantly affects the properties of substances like water and alcohols.
- **Unit Cells** are the smallest repeating units of a crystal lattice. They define the three-dimensional structure of the solid.
- **Packing Efficiency** refers to the fraction of space occupied by atoms in a unit cell. Common types include simple cubic, body-centered cubic, and face-centered cubic.
- **Polymers** are long-chain molecules that can form solid materials. Their properties depend on the nature of the repeating units and the strength of intermolecular forces.

Types of Chemical Calculations

Key to Answer on the Exam

- **Stoichiometry**: Calculations involving the relative amounts and masses of substances in a chemical reaction.
- **Thermochemistry**: Calculations involving heat energy changes during chemical processes.
- **Equilibrium**: Calculations involving the concentrations of reactants and products at equilibrium.
- **Acid-Base**: Calculations involving pH, pOH, and the concentrations of acids and bases.
- **Solubility**: Calculations involving the concentrations of ions in saturated solutions.
- **Electrochemistry**: Calculations involving the relationship between electrical energy and chemical reactions.
- **Gas Laws**: Calculations involving the pressure, volume, and temperature of gases.

Periodic Table & Trends

The periodic table is color-coded to show trends. Electronegativity increases from bottom-left to top-right. Ionization energy follows a similar trend. Atomic radius increases from top-left to bottom-right.

- **Electronegativity** is the ability of an atom to attract electrons in a bond. It increases across a period and decreases down a group.
- **Ionization Energy** is the energy required to remove an electron from a gaseous atom. It increases across a period and decreases down a group.
- **Atomic Radius** is the size of an atom. It increases down a group and across a period from right to left.
- **Electron Affinity** is the energy change when an electron is added to a gaseous atom. It generally increases across a period.
- **Metals** are located on the left side of the periodic table. They are typically shiny, malleable, and good conductors of electricity.
- **Nonmetals** are located on the right side of the periodic table. They are typically brittle and poor conductors of electricity.

Electronegativity Table

Electronegativity Trends:

- Increases from bottom-left to top-right.
- Fluorine (F) is the most electronegative element (4.0).
- Francium (Fr) is the least electronegative element (0.7).

The table shows electronegativity values for various elements, ranging from 0.7 for Francium to 4.0 for Fluorine.

Acids & Bases

Acids are substances that donate protons (H^+). They have a sour taste and turn blue litmus paper red. **Bases** are substances that accept protons. They have a bitter taste and turn red litmus paper blue. The **pH scale** measures the concentration of H^+ ions, ranging from 0 to 14.

Equilibrium & Solubility

Chemical Equilibrium is a state where the forward and reverse reactions occur at equal rates. The **Equilibrium Constant (K)** is a measure of the extent of a reaction. **Solubility** is the maximum amount of a substance that can dissolve in a given amount of solvent at a specific temperature.

Compound	Solubility (g/100 mL water at 25°C)
NaCl	36.0
KCl	34.5
CaCl ₂	147.0
MgCl ₂	54.9
AgCl	0.0001925
PbCl ₂	0.99

Le Chatelier's Principle states that if a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system will shift to counteract the change. **Common Ion Effect** occurs when a common ion is added to a solution, decreasing the solubility of a salt.



Redox Reactions involve the transfer of electrons between species. **Oxidation** is the loss of electrons, and **reduction** is the gain of electrons. The **Standard Reduction Potentials** are used to predict the spontaneity of redox reactions. The **Nernst Equation** relates the cell potential to the concentrations of the reactants and products.

Basic/Weak Substances

Weak Acids only partially dissociate in water. The **Acid Dissociation Constant (K_a)** measures the strength of a weak acid. **Weak Bases** also partially dissociate. The **Base Ionization Constant (K_b)** measures the strength of a weak base. **Buffers** are solutions that resist changes in pH when small amounts of acid or base are added.

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Wenbin Ji



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appear on this AP exam The Comparative Review provides a thorough discussion of the material most often tested on the Comparative exam Special emphasis is placed on the governments and politics of Britain France The former Soviet Union The People's Republic of China A glossary for the Comparative Government and Politics exam enables you to brush up on terms that you are likely to encounter on this test Scoring the Exam After the AP administrations more than 1 700 college professors and secondary school teachers are brought together to grade the exams during the first two weeks of June These readers are chosen from around the United States for their familiarity with the AP program The Multiple Choice sections of the Comparative Government if the number is less than zero enter zero Scoring the Free Response Section The following guide explains typical free response scoring criteria Score Explanation of Score 8 9 The thesis is extremely well developed and is supported with concrete evidence all aspects of the question have been addressed thoroughly discussions presented are balanced 6 7 The thesis is defined and supported the evidence provided is very organized the essay may be slightly imbalanced with one strong argument and one weak argument and or discuss one topic more thoroughly than the next sporadic factual errors may appear 5 A basic argument or thesis is provided evidence given supports the argument or thesis but does not clearly connect with the argument or thesis only the formal facets of the question are dealt with and informal facets are not adequately covered not all aspects of the question are discussed 4 The thesis is not organized and is not referred to in the essay the essay is little more than a recounting of facts and events the essay may be overloaded with data only one facet of the questions may be discussed numerous factual errors may appear 3 The thesis is weak evidence provided in support does not apply to the thesis factual errors are apparent 2 The thesis is very weak little or no factual evidence is provided to support the thesis irrelevant and inaccurate information appears 1 An attempt is made to answer the question but the support given is insignificant and the coverage of topics is incomplete 0 The question is not answered with any significance Free Response Part II Comparative only Score Explanation of Score 5 The thesis is extremely well developed and is supported with concrete evidence all aspects of the question have been addressed thoroughly discussions are presented in a balanced way 4 The thesis is defined and supported the evidence provided is very organized the essay may be slightly imbalanced with one strong argument and one weak argument likewise one topic may be more thoroughly explored than another may be marred by sporadic factual errors 3 A basic argument or thesis is presented evidence given supports the argument or thesis but does not clearly connect with the argument or thesis only the formal facets of the question are dealt with and informal facets are not adequately covered not all aspects of the question are discussed 2 The thesis is weak evidence provided in support does not apply to the thesis factual errors are apparent 1 An attempt is made to answer the question but the support given is insignificant and the coverage of topics is incomplete 0 The question is not answered with any significance It would be extremely helpful to find someone who is willing to score your essay your teachers or anyone who is familiar with the test material If you do ask the person to assign each of your U S and Comparative Part I essays a

score of 0 to 9 For your Comparative Part II essays use the 0 to 5 scale If you must grade your own essays try to be objective In addition you may want to give your essays three different grades For instance if you feel you did well try giving the essay a score of 5 6 or 7 to represent the various scores you may receive By underestimating what your score may be you are more likely to receive a better score on the actual exam Use the following formulae to determine your raw score for the Free Response section United States Exam Free Response Response 1 score \times 1 66 raw score Response 2 score \times 1 66 raw score Response 3 score \times 1 66 raw score Response 4 score \times 1 66 raw score Comparative Exam Response 1 score \times 1 66 raw score Response 2 score \times 1 66 raw score Response 3 score \times 3 raw score Response 4 score \times 3 raw score The Composite Score Once you have obtained your raw scores for both the Multiple Choice and the Free Response sections add the scores together to get your composite score United States Exam Multiple Choice raw score Free Response raw score composite score round to nearest whole number Score Essay 1 Score Essay 2 Score Essay 3 Score Essay 4 raw score Comparative Exam Multiple Choice raw score Free Response raw score composite score round to nearest whole number Now compare your composite score with the scale below Composite Score AP Grade 88 120 5 74 87 4 54 73 3 35 53 2 0 34 1 AP grades are interpreted as follows 5 extremely well qualified 4 well qualified 3 qualified 2 possibly qualified and 1 no recommendation Scores that Earn College Credit and or Advanced Placement Most colleges grant students who earn a 3 or above college credit and or advanced placement You should check with your school guidance office about specific college requirements Studying for Your AP Examination It is never too early to start studying The earlier you begin the more time you will have to sharpen your skills Do not procrastinate Cramming is not an effective way to study since it does not allow you the time needed to learn the test material It is very important for you to choose the time and place for studying that works best for you Some students may set aside a certain number of hours every morning to study while others may choose to study at night before going to sleep Other students may study during the day while waiting on a line or even while eating lunch Only you can determine when and where your study time will be most effective But be consistent and use your time wisely Work out a study routine and stick to it When you take the practice exam s try to make your testing conditions as much like the actual test as possible Turn your television and radio off and sit down at a quiet table free from distraction Make sure to time yourself As you complete the practice test s score your test s and thoroughly review the explanations to the questions you answered incorrectly but do not review too much during any one sitting Concentrate on one problem area at a time by reviewing the question and explanation and by studying our review s until you are confident that you completely understand the material Since you will be allowed to write in your test booklet during the actual exam you may want to write in the margins and spaces of this book when practicing However do not make miscellaneous notes on your answer sheet Mark your answers clearly and make sure the answer you have chosen corresponds to the question you are answering Keep track of your scores By doing so you will be able to gauge your progress and discover general weaknesses in particular sections You

should carefully study the reviews that cover the topics causing you difficulty as this will build your skills in those areas To get the most out of your studying time we recommend that you follow the Study Schedule which corresponds to the exam you are taking It details how you can best budget your time If you are taking both exams do not try to study for each at the same time Try alternating days by studying for the United States exam one day and the Comparative exam the next Test Taking Tips Although you may be unfamiliar with tests such as the Advanced Placement exams there are many ways to acquaint yourself with this type of examination and help alleviate your test taking anxieties Listed below are ways to help yourself become accustomed to the AP exam some of which may also be applied to other standardized tests Become comfortable with the format of the AP Examination in Government and Politics that you are taking When you are practicing to take the exam s simulate the conditions under which you will be taking the actual test s You should practice under the same time constraints as well Stay calm and pace yourself After simulating the test only a couple of times you will boost your chances of doing well and you will be able to sit down for the actual test much more confidently Know the directions and format for each section of the exam Familiarizing yourself with the directions and format of the different test sections will not only save you time but will also ensure that you are familiar enough with the AP exam to avoid nervousness and the mistakes caused by being nervous Work on the easier questions first If you find yourself working too long on one question make a mark next to it in your test booklet and continue After you have answered all of the questions that you can go back to the ones you have skipped Use the process of elimination when you are unsure of an answer If you can eliminate three of the answer choices you have given yourself a fifty fifty chance of getting the item correct since there will only be two choices left from which to make a guess If you cannot eliminate at least three of the answer choices you may choose not to guess as you will be penalized one quarter of a point for every incorrect answer Questions not answered will not be counted Be sure that you are marking your answer in the circle that corresponds to the number of the question in the test booklet Since the multiple choice section is graded by machine marking the wrong answer will throw off your score

African American Advanced Placement Chemistry Students and Their Developing Study Habits Natalie D. Rasmussen, 2006

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course description outline and sample AP test questions It covers only the information tested on the new exam so you can make the most of your valuable study time Our targeted review focuses on the Big Ideas that will be covered on the exam Explanations of the AP Chemistry Labs are also included Expert Test taking Strategies This Crash Course presents detailed question level strategies for answering both the multiple choice and essay questions By following this advice you can boost your score in every section of the test Take REA's Online Practice Exam After studying the material in the Crash Course go to the online REA Study Center and test what you've learned Our practice exam features timed testing detailed explanations of answers and automatic scoring analysis The exam is balanced to include every topic and type of question found on the actual AP exam so you know you're studying the smart way Whether you're cramming for the test at the last minute looking for extra review or want to study on your own in preparation for the exams this is the study guide every AP Chemistry student must have When it's crucial crunch time and your Advanced Placement exam is just around the corner you need REA's Crash Course for AP Chemistry About the Author Adrian Dingle is a chemistry educator and author with 24 years of experience teaching in the United States and the United Kingdom He is the creator of the award winning chemistry website www.adriandingleschemistrypages.com The focus of Mr Dingle's teaching career has been on preparing students for standardized tests AP and SAT tests in the United States GCSE's and A levels in the United Kingdom and International Baccalaureate in both countries An Englishman he lives in Atlanta Georgia where he teaches at The Westminster Schools He holds a B Sc Hons Chemistry and a Postgraduate Certificate in Education both from the University of Exeter in England In addition to writing this Crash Course Mr Dingle has written *The Periodic Table Elements With Style* *How To Make A Universe With 92 Ingredients* and *SAT Chemistry Crash Course* He is the 2011 winner of the School Library Association of the UK's Information Book Award and in 2012 was honored with the prestigious literary prize *Wissenschaftsbuch des Jahres* sponsored by the Austrian Ministry of Science and Research **New Horizons in Mathematics and Science Education**, 2001

ENC Focus, 2001 *The American Educational Review*, 1912

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