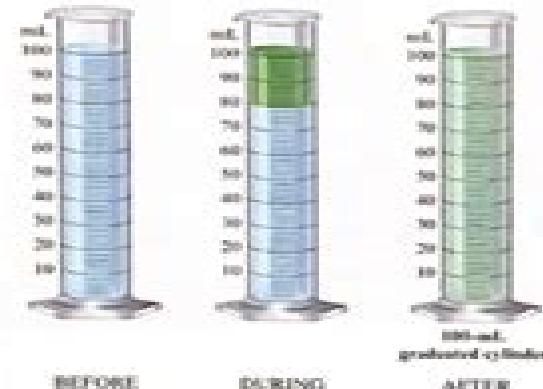


Multiple Choice: (from Zumdahl Chemistry, 9th Edition, AP Edition)

Refer to the following diagram for questions 1 – 3.

A 100.0 mL Graduated Cylinder is filled with 100.0 mL of water. A few drops of food colouring are added to the top of the water, without mixing. Initially, the top of the graduated cylinder appears to be very dark, while the bottom remains clear and colourless. After 5 min, the liquid in the graduated cylinder appears to be uniformly coloured.



1. Classify the contents in the graduated cylinder before any food colouring is added.

(A) Element (B) Compound. (C) Heterogeneous mixture
(D) Homogeneous mixture

2. Classify the contents in the graduated cylinder just after the food colouring is added.

(A) Element (B) Compound. (C) Heterogeneous Mixture (D) Homogeneous mixture

3. Classify the contents in the graduated cylinder after the 5 minutes has elapsed.

(A) Element (B) Compound. (C) Heterogeneous Mixture (D) Homogeneous mixture

4. A 10.00 g piece of metal is submerged in a graduated cylinder initially containing 20.00 mL of water. With the metal in the cylinder, the level of water is recorded as 21.40 mL. Determine the identity of the metal.

(A) Al ($\rho = 2.70 \text{ g/mL}$)
(B) Zn ($\rho = 7.13 \text{ g/mL}$)
(C) Ag ($\rho = 10.49 \text{ g/mL}$)
(D) Au ($\rho = 19.32 \text{ g/mL}$)

5. $^{240}_{94}\text{Pu} \rightarrow X + \alpha$

In the decay represented above, what is the nucleon number of particle X?

(A) 92 (B) 96 (C) 235 (D) 243

6. Which of the following statements about atoms is CORRECT?

(A) Atoms are charged because they have different numbers of protons and neutrons.
(B) All atoms of an element have the same number of nucleons and electrons.
(C) An atom is mostly filled with protons.
(D) The nucleus is negatively charged.

6. When each of the following is heated to 50°C and the temperature is held constant for 5 min, which one undergoes only a physical change.

(A) egg (B) steak (C) ice cream (D) cake batter

7. A sample of Ti was ionized by electron impact in a mass spectrometer. Information from the mass spectrum about the isotopes of Ti in the sample is shown in Table 2.

Calculate the relative atomic mass of Ti in this sample. Report answers to one decimal place.

Table 1. Results from mass spectrum of Ti sample

<i>m/z</i>	46	47	48	49
Percent Abundance	9.1	7.8	74.6	8.5

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