

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and  $K_a$  or  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                  |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain                | 6. acid rain     |
|                   | 7. acid solution            | 8. pH            |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the unknown quantity that are indicated in the problem. Be sure to include units whenever a unit is appropriate or required.

1. What substance is \_\_\_\_\_, a weak hydrogen ion donor?
2. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions present in a solution and equal to \_\_\_\_\_.
3. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
4. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
5. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
6. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
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12. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
13. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
14. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
15. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
16. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
17. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
18. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
19. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.
20. \_\_\_\_\_ is the concentration of  $\text{H}^+$  ions in a solution.

1. \_\_\_\_\_
2. \_\_\_\_\_
3. \_\_\_\_\_
4. \_\_\_\_\_
5. \_\_\_\_\_
6. \_\_\_\_\_
7. \_\_\_\_\_
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20. \_\_\_\_\_

# 19 2 Hydrogen Ions And Acidity Answer Key

**Barbara Janson Cohen, Kerry L. Hull**



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