

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonds

There are two types of bonds: ionic bonds and covalent bonds.

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Strength of Bonds



Bond and Bond Energy



Bond Energy



There are two types of bonds

Definition of Polarity



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

A difference of 1.7 or higher indicates that the electrons are not shared (ionic bond).

A difference of 1.7 or higher indicates that the electrons are not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector dipole moment is a product of the charge difference and the distance between the charges.
- $\mu = q \times r$ (where q is the charge difference and r is the distance between the charges).

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments of the bonds.



Sum the vector of the dipole moments to find the net dipole.

Bond Length and Bond Strength

The length of the bond is inversely proportional to the strength of the bond.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds.

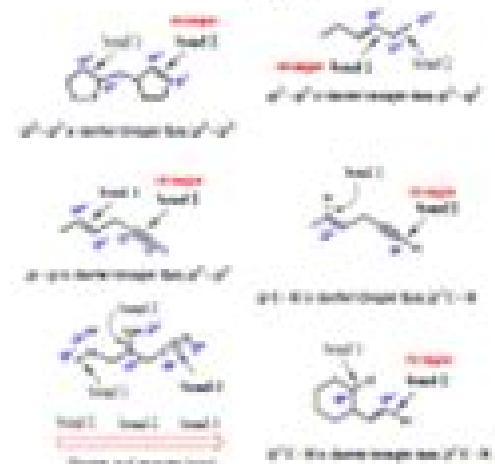


The length of the bond is inversely proportional to the strength of the bond.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	460 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
Length	147 pm	133 pm	121 pm
Strength	356 kJ/mol	460 kJ/mol	591 kJ/mol

Higher s character means stronger bond

Examples



Best Organic Chemistry Study Guide

**T. W. Graham Solomons, Craig B.
Fryhle, Scott A. Snyder**



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Further Explorations are short discussions that provide additional depth on content in the text Reaction Reviews present the reactions in the chapter in detailed outline form This summary consolidates the reactions of the chapter in one place and presents not only the reactions themselves but also the essential features of their mechanisms and stereochemistry for cases in which these issues are discussed in the text Solutions are provided for all of the over 1600 problems in the text Many include worked out examples in which not only the answer is given but also the detailed approach to the solution is discussed

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