

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. K_a | 9. K_b |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constants indicated that are associated with the reaction. Be sure to include the correct number of significant figures.

1. Write the equation for the autoionization of water.
2. Write the equation for the autoionization of water at 25°C.
3. Write the equation for the autoionization of water at 25°C.
4. Write the equation for the autoionization of water at 25°C.
5. Write the equation for the autoionization of water at 25°C.
6. Write the equation for the autoionization of water at 25°C.
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15. Write the equation for the autoionization of water at 25°C.
16. Write the equation for the autoionization of water at 25°C.
17. Write the equation for the autoionization of water at 25°C.
18. Write the equation for the autoionization of water at 25°C.
19. Write the equation for the autoionization of water at 25°C.
20. Write the equation for the autoionization of water at 25°C.

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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20. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

19 2 Hydrogen Ions And Acidity Work Sheet Answers



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