

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentration from pH and hydroxide ion concentration from pOH.
3. Calculate the pH of a weak acid.

## Vocabulary

- |                   |                               |                       |
|-------------------|-------------------------------|-----------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution      |
| 4. autoionization | 5. acid rain                  | 6. acid rain solution |
|                   | 7. acid solution              | 8. pH                 |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Read the paragraph and use the information to determine if the sentence is true or false. Write true or false in the space provided. If the sentence is false, rewrite it so that it is true. Do not copy the sentence.

1. When a substance has a pH of 1, it is an acid solution.
2. The concentration of  $\text{H}^+$  in a solution of 0.001 M is 0.001 M.
3. The pH of a solution of 0.001 M is 3.
4. The pH of a solution of 0.001 M is 3.
5. The pH of a solution of 0.001 M is 3.
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# 19 2 Hydrogen Ions And Acidity Answer Key

**Frank R. Foulkes**



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