

**Chemical Reactions:**

1. The following reaction  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  is given. What is the  $\Delta H^\circ_f$  for the reaction?

A) greater than 2.00 kcal/mol  
 B) greater than 1.00 kcal/mol  
 C) less than 1.00 kcal/mol  
 D) less than 2.00 kcal/mol  
 E) greater than 4.00 kcal/mol

2. A 1.0-mol sample of  $\text{PCl}_3$  reacts with excess oxygen according to the equation:  $2\text{PCl}_3(\text{g}) \rightarrow 2\text{PCl}_5(\text{g}) + \text{O}_2(\text{g})$ . Below, which reaction will yield the greatest amount of  $\text{O}_2$  after  $\text{PCl}_3$  is removed?

A) 0.0 mol  
 B) 1.0 mol  
 C) 1.2 mol  
 D) 1.7 mol  
 E) 2.0 mol

3. The reaction of  $\text{NO}_2$  is step characterized by the equation:  $4\text{NO}_2 \rightarrow 3\text{O}_2 + 2\text{N}_2\text{O}_5$ . If 1.0 mol of  $\text{NO}_2$  is added to 1.0 mol of  $\text{N}_2\text{O}_5$  and the reaction is allowed to proceed, how will the reaction rate change?

A) 1.000 mol/d  
 B) 0.750 mol/d  
 C) 1.000 mol/d  
 D) 1.50 mol/d  
 E) 2.00 mol/d

4. Calculated are the following reactions, or from the experiments, are unknowns, calculated.

$\text{H}_2\text{O}_2(\text{l}) + \text{H}_2\text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O}_2(\text{l}) = 300\text{J/g}$

$\text{H}_2 + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$  (enthalpy of difference,  $\Delta H^\circ_f$  for production of 1.00 mol of  $\text{H}_2\text{O}_2$ , which is the difference of 200)

A)  $\text{O}_2$   
 B)  $\text{H}_2$   
 C)  $\text{H}_2$   
 D)  $\text{O}_2$   
 E) C

5. Refer to the following reaction:

$\text{NO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + \text{O}_2(\text{g})$

Major steady-state levels of  $\text{NO}_2$  are given below. In which behavior is each?

A)  $1.2 \times 10^{-4}$   
 B)  $1.25 \times 10^{-4}$   
 C)  $1.30 \times 10^{-4}$   
 D)  $1.35 \times 10^{-4}$   
 E) Insufficient information

6. The following reaction  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  is given. Below, which reaction will yield the greatest amount of  $\text{O}_2$  after 1.00 mol of hydrogen and oxygen react?

A) 1.00 mol  
 B) 1.20 mol  
 C) 1.25 mol  
 D) 1.30 mol  
 E) 1.35 mol

7. Refer to the following reaction:

$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{OH} + \text{H}_2\text{O}$

What mass of oxygen,  $\text{O}_2$ , is required to react completely with 1.00 g of  $\text{C}_2\text{H}_6$ ?

A)  $1.00 \times 10^{-2}$  g  
 B)  $1.00 \times 10^{-1}$  g  
 C)  $1.00 \times 10^{-3}$  g  
 D)  $1.00 \times 10^{-4}$  g  
 E)  $1.00 \times 10^{-5}$  g

8. Refer reaction neutralization and synthesis of ammonia. In the following reaction, which is the limiting reactant?



A) 1.00  $\times 10^{-2}$  mol/deciliter  
 B) 1.00  $\times 10^{-3}$  mol/deciliter  
 C)  $1.00 \times 10^{-4}$  mol/deciliter  
 D)  $1.00 \times 10^{-5}$  mol/deciliter  
 E)  $1.00 \times 10^{-6}$  mol/deciliter

9. Calculate the mass of carbon dioxide produced from 1.00 g of octane,  $\text{C}_8\text{H}_{18}$ , in the following reaction:



A)  $1.75 \times 10^{-2}$  g  
 B)  $1.75 \times 10^{-1}$  g  
 C)  $1.75 \times 10^{-3}$  g  
 D)  $1.75 \times 10^{-4}$  g  
 E)  $1.75 \times 10^{-5}$  g

10. Calculate the mass of oxygen required to react with 1.00 g of sulfur in the following reaction:



A)  $1.00 \times 10^{-2}$  mol/deciliter  $\text{O}_2$   
 B)  $1.00 \times 10^{-3}$  mol/deciliter  $\text{O}_2$   
 C)  $1.00 \times 10^{-4}$  mol/deciliter  $\text{O}_2$   
 D)  $1.00 \times 10^{-5}$  mol/deciliter  $\text{O}_2$   
 E)  $1.00 \times 10^{-6}$  mol/deciliter  $\text{O}_2$

# Multiple Choice Stoichiometry Test

**Martha Mackin**

## Multiple Choice Stoichiometry Test:

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