

# Chapter 1

## Problem Solutions

1.1

(a) fcc: 8 corner atoms  $\times 1/8 = 1$  atom  
       6 face atoms  $\times 1/2 = 3$  atoms  
       Total of 4 atoms per unit cell

(b) bcc: 8 corner atoms  $\times 1/8 = 1$  atom  
       1 enclosed atom  $= 1$  atom  
       Total of 2 atoms per unit cell

(c) Diamond: 8 corner atoms  $\times 1/8 = 1$  atom  
       6 face atoms  $\times 1/2 = 3$  atoms  
       4 enclosed atoms  $= 4$  atoms  
       Total of 8 atoms per unit cell

1.2

(a) 4 Ga atoms per unit cell

$$\text{Density} = \frac{4}{(5.65 \times 10^{-8})^3} \Rightarrow$$

Density of Ga =  $2.22 \times 10^{27} \text{ cm}^{-3}$

4 As atoms per unit cell, so that

Density of As =  $2.22 \times 10^{27} \text{ cm}^{-3}$

(b)

8 Ge atoms per unit cell

$$\text{Density} = \frac{8}{(5.65 \times 10^{-8})^3} \Rightarrow$$

Density of Ge =  $4.44 \times 10^{27} \text{ cm}^{-3}$

1.3

(a) Simple cubic lattice,  $\alpha = 2r$

Unit cell vol =  $\alpha^3 = (2r)^3 = 8r^3$

1 atom per cell, so atom vol =  $(1) \left( \frac{4\pi r^3}{3} \right)$

Then

$$\text{Ratio} = \frac{\left( \frac{4\pi r^3}{3} \right)}{8r^3} \times 100\% \Rightarrow \text{Ratio} = 52.4\%$$

(b) Face-centered cubic lattice

$$d = 4r = \alpha\sqrt{2} \Rightarrow \alpha = \frac{d}{\sqrt{2}} = 2\sqrt{2}r$$

Unit cell vol =  $\alpha^3 = (2\sqrt{2}r)^3 = 16\sqrt{2}r^3$

4 atoms per cell, so atom vol =  $4 \left( \frac{4\pi r^3}{3} \right)$

Then

$$\text{Ratio} = \frac{4 \left( \frac{4\pi r^3}{3} \right)}{16\sqrt{2}r^3} \times 100\% \Rightarrow \text{Ratio} = 74\%$$

(c) Body-centered cubic lattice

$$d = 4r = \alpha\sqrt{3} \Rightarrow \alpha = \frac{4}{\sqrt{3}}r$$

Unit cell vol =  $\alpha^3 = \left( \frac{4}{\sqrt{3}}r \right)^3$

2 atoms per cell, so atom vol =  $2 \left( \frac{4\pi r^3}{3} \right)$

Then

$$\text{Ratio} = \frac{2 \left( \frac{4\pi r^3}{3} \right)}{\left( \frac{4r}{\sqrt{3}} \right)^3} \times 100\% \Rightarrow \text{Ratio} = 68\%$$

(d) Diamond lattice

$$\text{Body diagonal} = d = 8r = \alpha\sqrt{3} \Rightarrow \alpha = \frac{8}{\sqrt{3}}r$$

Unit cell vol =  $\alpha^3 = \left( \frac{8r}{\sqrt{3}} \right)^3$

8 atoms per cell, so atom vol =  $8 \left( \frac{4\pi r^3}{3} \right)$

Then

$$\text{Ratio} = \frac{8 \left( \frac{4\pi r^3}{3} \right)}{\left( \frac{8r}{\sqrt{3}} \right)^3} \times 100\% \Rightarrow \text{Ratio} = 34\%$$

1.4

From Problem 1.3, percent volume of fcc atoms is 74%; Therefore after evaporation, Volume =  $0.74 \text{ cm}^3$

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**S Ashworth**

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