

Chemical Reactions:

1. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. What is the ΔH°_f for the reaction?

A) greater than 2.00 kcal/mol
 B) greater than 1.00 kcal/mol
 C) less than 1.00 kcal/mol
 D) less than 2.00 kcal/mol
 E) greater than 4.00 kcal/mol

2. A 1.0-mol sample of H_2CO_3 reacts with excess CaCO_3 according to the equation: $\text{CaCO}_3(s) + 2\text{H}_2\text{CO}_3(aq) \rightarrow \text{CaHCO}_3(s) + 2\text{H}_2\text{O}(g)$. Below, which samples of O_2 are formed in the reaction $\text{CaHCO}_3(s) \rightarrow \text{CaO}(s) + \text{CO}_2(g) + \text{H}_2\text{O}(g)$?

A) 0.0 mol
 B) 1.0 mol
 C) 1.2 mol
 D) 1.7 mol
 E) 2.0 mol

3. The reaction of H_2 with Cl_2 is represented by the equation: $2\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$. If 1.0 mol of H_2 reacts with 1.0 mol of Cl_2 , which of the following samples of HCl are formed? Assume that H_2 reacts completely.

A) 0.000 mol
 B) 0.700 mol
 C) 1.000 mol
 D) 1.500 mol
 E) 2.000 mol

4. Calculated are the following reactions, or both? (i) $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ and (ii) $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$.

What is the ΔH°_f of H_2O formed in reaction (i) if reaction (ii) is measured?

A) 0.000 kcal/mol
 B) 0.444 $\times 10^3$ kcal/mol
 C) 0.888 $\times 10^3$ kcal/mol
 D) 1.333 $\times 10^3$ kcal/mol
 E) 1.778 $\times 10^3$ kcal/mol

5. Calculated are the following reactions, or both? (i) $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ and (ii) $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$.

What is the ΔH°_f of CO_2 formed in reaction (ii) if reaction (i) is measured?

A) 0.000 kcal/mol
 B) 0.444 $\times 10^3$ kcal/mol
 C) 0.888 $\times 10^3$ kcal/mol
 D) 1.333 $\times 10^3$ kcal/mol
 E) 1.778 $\times 10^3$ kcal/mol

6. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. What is the ΔS° for the reaction?

A) 1.00 $\times 10^3$ J/mol-K
 B) 2.00 $\times 10^3$ J/mol-K
 C) 3.00 $\times 10^3$ J/mol-K
 D) 4.00 $\times 10^3$ J/mol-K
 E) 5.00 $\times 10^3$ J/mol-K

7. The following reaction $\text{CaHCO}_3(s) \rightarrow \text{CaO}(s) + \text{CO}_2(g) + \text{H}_2\text{O}(g)$ is measured. What is the ΔS° for the reaction?

A) 0.000 $\times 10^3$ J/mol-K
 B) 0.444 $\times 10^3$ J/mol-K
 C) 0.888 $\times 10^3$ J/mol-K
 D) 1.333 $\times 10^3$ J/mol-K
 E) 1.778 $\times 10^3$ J/mol-K

8. The following reaction $\text{CaHCO}_3(s) \rightarrow \text{CaO}(s) + \text{CO}_2(g) + \text{H}_2\text{O}(g)$ is measured. What is the ΔS° for the reaction?

A) 0.000 $\times 10^3$ J/mol-K
 B) 0.444 $\times 10^3$ J/mol-K
 C) 0.888 $\times 10^3$ J/mol-K
 D) 1.333 $\times 10^3$ J/mol-K
 E) 1.778 $\times 10^3$ J/mol-K

9. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. What is the ΔS° for the reaction?

A) 1.00 $\times 10^3$ J/mol-K
 B) 2.00 $\times 10^3$ J/mol-K
 C) 3.00 $\times 10^3$ J/mol-K
 D) 4.00 $\times 10^3$ J/mol-K
 E) 5.00 $\times 10^3$ J/mol-K

10. Calculate the following reaction enthalpy of reaction $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{OH} + \text{H}_2\text{O}$. What is the ΔH°_f of CH_3OH if C_2H_6 is measured to react completely with 1.00 g of O_2 ?

A) 0.000 $\times 10^3$ J/mol-K
 B) 0.444 $\times 10^3$ J/mol-K
 C) 0.888 $\times 10^3$ J/mol-K
 D) 1.333 $\times 10^3$ J/mol-K
 E) 1.778 $\times 10^3$ J/mol-K

11. Calculate the reaction enthalpy of reaction $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{OH} + \text{H}_2\text{O}$. What is the ΔH°_f of CH_3OH if C_2H_6 is measured to react completely with 1.00 g of O_2 ?

A) 0.000 $\times 10^3$ J/mol-K
 B) 0.444 $\times 10^3$ J/mol-K
 C) 0.888 $\times 10^3$ J/mol-K
 D) 1.333 $\times 10^3$ J/mol-K
 E) 1.778 $\times 10^3$ J/mol-K

12. Calculate the reaction enthalpy of reaction $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{OH} + \text{H}_2\text{O}$. What is the ΔH°_f of CH_3OH if C_2H_6 is measured to react completely with 1.00 g of O_2 ?

13. Calculate the reaction enthalpy of reaction $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{OH} + \text{H}_2\text{O}$. What is the ΔH°_f of CH_3OH if C_2H_6 is measured to react completely with 1.00 g of O_2 ?

A) 0.000 $\times 10^3$ J/mol-K
 B) 0.444 $\times 10^3$ J/mol-K
 C) 0.888 $\times 10^3$ J/mol-K
 D) 1.333 $\times 10^3$ J/mol-K
 E) 1.778 $\times 10^3$ J/mol-K

14. Calculate the reaction enthalpy of reaction $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{OH} + \text{H}_2\text{O}$. What is the ΔH°_f of CH_3OH if C_2H_6 is measured to react completely with 1.00 g of O_2 ?

15. Calculate the reaction enthalpy of reaction $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{OH} + \text{H}_2\text{O}$. What is the ΔH°_f of CH_3OH if C_2H_6 is measured to react completely with 1.00 g of O_2 ?

A) 0.000 $\times 10^3$ J/mol-K
 B) 0.444 $\times 10^3$ J/mol-K
 C) 0.888 $\times 10^3$ J/mol-K
 D) 1.333 $\times 10^3$ J/mol-K
 E) 1.778 $\times 10^3$ J/mol-K

Multiple Choice Stoichiometry Test

Roger Slade, Maureen Slade

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Three Cognitive Skills in Chemistry and Their Application to Stoichiometry Ardas Ozsogomonyan, 1977 **A Summary of Science Education**, 1988 Leopold E.

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