

Chemical Reactions:

1. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. What is the ΔH°_f for the reaction?

A) greater than 2.00 kcal/mol
 B) greater than 1.00 kcal/mol
 C) less than 1.00 kcal/mol
 D) less than 2.00 kcal/mol
 E) greater than 4.00 kcal/mol

2. A 1.0-mol sample of PCl_3 reacts with excess oxygen according to the equation: $2\text{PCl}_3(\text{g}) \rightarrow 2\text{PCl}_5(\text{g}) + \text{O}_2(\text{g})$. Below, which reaction will yield the greatest amount of O_2 after PCl_3 is reacted?

A) 0.0 mol
 B) 1.0 mol
 C) 1.2 mol
 D) 1.7 mol
 E) 2.0 mol

3. The reaction of NO_2 is step characterized by the equation: $4\text{NO}_2 \rightarrow 3\text{O}_2 + 2\text{N}_2\text{O}_5$. If 1.0 mol of NO_2 is reacted with 1.0 mol of O_2 , which will affect the rate of the reaction?

A) O_2 (0.001 mol)
 B) O_2 (0.002 mol)
 C) O_2 (0.003 mol)
 D) O_2 (0.004 mol)
 E) O_2 (0.005 mol)

4. Calculated the following reaction, ΔH°_f for the reactants are unknowns. $\text{H}_2\text{O}_2(\text{l}) + \text{H}_2\text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$

ΔH°_f for H_2O_2 is 20.0 kcal/mol. In which reaction is ΔH°_f for the reactants less than 20.0 kcal/mol?

A) A
 B) B
 C) C
 D) D
 E) E

5. Refer to the following reaction:

$$\text{NO}_2(\text{g}) + \text{NO}(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$$

What is the ΔH°_f for the reaction?

A) 1.00
 B) 1.00×10^3
 C) 1.00×10^4
 D) 1.00×10^5
 E) Insufficient information

6. The following reaction $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}$ is given. Below, which number of molecules would be present in 1.00 mol of hydrogen and oxygen mixture?

A) 1.00 mol
 B) 2.00 mol
 C) 0.500 mol
 D) 0.50 mol
 E) 0.250 mol

7. Refer to the following reaction:

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightleftharpoons \text{CH}_3\text{OH} + \text{H}_2\text{O}$$

What mass of oxygen, $\text{O}_2(\text{g})$ is required to react completely with 1.00 g of C_2H_6 ?

A) 0.001×10^3 g
 B) 0.01×10^3 g
 C) 0.005×10^3 g
 D) 0.015×10^3 g
 E) 0.0005×10^3 g

8. Refer to the reaction and questions of reaction 7. How many molecules of O_2 are required to react with 1.00 mol of C_2H_6 ?

$\text{C}_2\text{H}_6(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{CH}_3(\text{g}) + \text{H}_2\text{O}(\text{g})$

A) 1.00×10^{23} molecules
 B) 1.00×10^{24} molecules
 C) 4.00×10^{23} molecules
 D) 2.00×10^{23} molecules
 E) 0.50×10^{23} molecules

9. Calculate the mass of carbon dioxide produced from 1.00 g of octane, C_8H_{18} , in the following reaction:



10. Calculate the number of oxygen atoms required to react with 1.00 g of sulfur in the following reaction:



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Three Cognitive Skills in Chemistry and Their Application to Stoichiometry Ardas Ozsogomonyan, 1977 **A Summary of Science Education**, 1988 Leopold E.

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