

**Chemical Reactions:**

1. The following reaction  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  is given. What is the  $\Delta H^\circ_f$  for  $\text{O}_2$  in  $\text{kJ/mol}$ ?
   
A) greater than 2.0  $\text{kJ/mol}$   $\text{O}_2$   
 B) greater than 1.0  $\text{kJ/mol}$   $\text{O}_2$   
 C) equal to 1.0  $\text{kJ/mol}$   $\text{O}_2$   
 D) less than 1.0  $\text{kJ/mol}$   $\text{O}_2$   
 E) greater than 4.0  $\text{kJ/mol}$   $\text{O}_2$

2. A 1.0-molar sample of  $\text{H}_2\text{SO}_4$  reacts with excess  $\text{CaCO}_3$  according to the equation:  

$$2\text{H}_2\text{SO}_4(\text{aq}) \rightarrow 2\text{H}_2\text{O} + \text{SO}_2(\text{g})$$
 Below, which samples of  $\text{O}_2$  are consumed in the reaction?
   
A) 0.0  $\text{mol}$   
 B) 1.0  $\text{mol}$   
 C) 1.2  $\text{mol}$   
 D) 1.7  $\text{mol}$   
 E) 2.0  $\text{mol}$

3. The reaction of  $\text{NO}_2$  is step characterized by the equation:  $4\text{NO}_2 \rightarrow 3\text{O}_2 + 2\text{N}_2\text{O}_5$ . If 1.0 mol of  $\text{NO}_2$  reacts with 1.0 mol of  $\text{O}_2$ , which of the following is true after the reaction has reached completion?
   
A) 1.0  $\text{mol}$   
 B) 1.75  $\text{mol}$   
 C) 1.80  $\text{mol}$   
 D) 1.94  $\text{mol}$   
 E) 2.0  $\text{mol}$

4. Calculated are the following reactions, or both?
   
Requirements are molarities, not amounts.  
 $\text{H}_2\text{O}_2 + \text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O}_2\text{H}_2\text{O} \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ 
  
If 1.0  $\text{g}$  of  $\text{O}_2$  is added to excess  $\text{H}_2\text{O}_2$ , what is the total amount of  $\text{O}_2$  produced in 1.0  $\text{mol}$  of  $\text{H}_2\text{O}_2$ , which is the following?
   
A) 0.0  
 B) 0.40  
 C) 0.60  
 D) 1.0  
 E) 1.2

5. Under the following conditions,  $\text{NO}_2 + \text{NO}_2 \rightarrow 2\text{NO} + \text{O}_2$  is:
   
Major steady-state levels of  $\text{NO}_2$  are given below. In which behavior is each?
   
A) 1.0  $\text{kJ}$   
 B)  $1.00 \times 10^2$   
 C)  $1.00 \times 10^3$   
 D)  $1.00 \times 10^4$   
 E) Insufficient information

6. The following reaction  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  is:
   
Under which steady-state conditions would the greatest amount of  $\text{O}_2$  (mol) be produced?
   
A) 1.0  $\text{mol}$   
 B) 2.0  $\text{mol}$   
 C) 0.50  $\text{mol}$   
 D) 0.10  $\text{mol}$   
 E) 0.0001  $\text{mol}$

7. Under the following conditions,  $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  is:
   
What mass of oxygen,  $\text{O}_2$ , is required to react completely with 1.0  $\text{g}$  of  $\text{CH}_4$ ?
   
A)  $1.00 \times 10^{-2}$   $\text{g}$   
 B)  $1.00 \times 10^{-1}$   $\text{g}$   
 C)  $1.00 \times 10^0$   $\text{g}$   
 D)  $1.00 \times 10^1$   $\text{g}$   
 E)  $1.00 \times 10^2$   $\text{g}$

8. Under steady-state conditions, what is the total amount of  $\text{O}_2$  (mol) produced in the following reaction?
 
$$\text{C}_2\text{H}_6 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$$
 A)  $1.00 \times 10^{12}$  mol/second  
 B)  $1.00 \times 10^{13}$  mol/second  
 C)  $1.00 \times 10^{14}$  mol/second  
 D)  $1.00 \times 10^{15}$  mol/second  
 E)  $1.00 \times 10^{16}$  mol/second

9. Calculate the mass of carbon dioxide produced from 1.0  $\text{g}$  of octane,  $\text{C}_8\text{H}_{18}$ , in the following reaction:
 
$$2\text{C}_8\text{H}_{18} + 25\text{O}_2 \rightarrow 16\text{CO}_2 + 18\text{H}_2\text{O}$$

A)  $1.78 \times 10^{-2}$   $\text{g}$   
 B)  $1.78 \times 10^{-1}$   $\text{g}$   
 C)  $1.78 \times 10^0$   $\text{g}$   
 D)  $1.78 \times 10^1$   $\text{g}$   
 E)  $1.78 \times 10^2$   $\text{g}$

10. Calculate the molarities of oxygen required to react with 1.0  $\text{g}$  of sulfur in the following reaction:
 
$$2\text{S} + 3\text{O}_2 \rightarrow 2\text{SO}_2$$

A)  $1.00 \times 10^{12}$  mol/second  $\text{O}_2$   
 B)  $1.00 \times 10^{13}$  mol/second  $\text{O}_2$   
 C)  $1.00 \times 10^{14}$  mol/second  $\text{O}_2$   
 D)  $1.00 \times 10^{15}$  mol/second  $\text{O}_2$   
 E)  $1.00 \times 10^{16}$  mol/second  $\text{O}_2$

# Multiple Choice Stoichiometry Test

**Ardas Ozsogomonyan**

## Multiple Choice Stoichiometry Test:

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*Journal of Chemical Education*, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers *Tests* Richard C. Sweetland, Daniel J. Keyser, 1983 Descriptions of over 3000 tests in English intended as a guide for psychologists educators and other personnel who need test information to meet their assessment needs Topical arrangement under 3 main sections of psychology education and business Each entry gives test name associated personal names intended age group purpose description time range scoring cost and publisher Indexes by test titles authors publishers visually impaired tests and scoring services *The ETS Test Collection Catalog* Educational Testing Service. Test Collection, 1993 The major source of information on the availability of standardized tests Wilson Library Bulletin Covers commercially available standardized tests and hard to locate

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- Chapter 2: Essential Elements of Multiple Choice Stoichiometry Test
- Chapter 3: Multiple Choice Stoichiometry Test in Everyday Life
- Chapter 4: Multiple Choice Stoichiometry Test in Specific Contexts
- Chapter 5: Conclusion

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