

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. K_a | 6. K_b |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the K_a or K_b for the acid or base in parentheses. (100 points)

1. HCl (strong acid)
2. H_2SO_4 (strong acid)
3. HNO_3 (strong acid)
4. H_2SO_4 (strong acid)
5. H_2SO_4 (strong acid)
6. H_2SO_4 (strong acid)
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19. H_2SO_4 (strong acid)
20. H_2SO_4 (strong acid)

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20. H_2SO_4 (strong acid)

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Cecie Starr, Ralph Taggart



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