

Name: \_\_\_\_\_

Date: \_\_\_\_\_

## Behavior of Gases Guided Notes

### Pressure

- \_\_\_\_\_ is the amount of force exerted per unit of area, or  $P = F/A$ .
- A balloon and a bicycle tire are considered to be \_\_\_\_\_.
- They remain inflated because of \_\_\_\_\_ the air particles have with the walls of their container.

### Pressure

- This collection of forces, caused by the collisions of the particles, pushes the walls of the container \_\_\_\_\_.
- If more air is pumped into the balloon, the number of air \_\_\_\_\_ is increased.
- This causes more collisions with the walls of the container, which causes it to expand.

### Pressure

Pressure is measured in a unit called \_\_\_\_\_ (Pa), the SI unit of pressure.

### Pressure

At sea level, atmospheric pressure is 101.3 kPa.

At Earth's surface, the atmosphere exerts a force of about 101,300 N on every \_\_\_\_\_ —about the weight of a large truck.

### Boyle's Law

What happens to the gas pressure if you decrease the size of the container?

If you squeeze gas into a smaller space, its particles will strike the walls \_\_\_\_\_ often —giving an increased pressure. The opposite is true, too.

### Boyle's Law

- Robert Boyle (1627-1691), a British scientist, described this property of gases.
- According to **Boyle's law**, if you \_\_\_\_\_ the volume of a container of gas and hold the temperature constant, the pressure of the gas will \_\_\_\_\_.

# Behavior Of Gases Guided Practice Problems Answers

**Matt Kingsley**



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