

Honors Final Review

*Note: this is in no way guaranteed to be comprehensive. Anything from the year is fair game, but the emphasis will be on more recent material. **Bolded** items are most important.*

Chapter 1: Introductory stuff.

- Qualitative vs quantitative measurements.
- Scientific method

Chapter 2: Atoms, elements, and mixtures.

- Matter is made of up atoms
- **Elements, compounds, and molecules:** elements are only one type of atom, compounds have multiple elements, and molecules are atoms that are stuck together.
- States of matter—solid, liquid, and gas—and their properties.
- Physical versus chemical changes.
- Mixtures: homogeneous and heterogeneous.

Chapter 3: The elements and ions.

- Dalton's atomic theory: elements are made of atoms, all atoms of a given element are identical, atoms of different elements are different, and reactions don't alter the atoms—only rearrange them.
- Structure of the atom: protons and neutrons in the nucleus, electrons around the outside.
- J.J. Thomson discovered electrons.
- Rutherford shot particles at gold foil and saw some scatter, some go through, meaning that part of the atom was hard while most was empty space.
- Element symbols: atomic number and atomic mass
- **Atomic number = number of protons, defines an element.**
- **Atomic mass = # protons + # neutrons**
- Isotopes have same number of protons, different number of neutrons (same element, but weigh different amounts).
- **Number of electrons changes, leading to charges.** Things with a charge are called ions.
- Positively charged atoms/groups of atoms are cations. Negatively charged ones are anions.
- Periodic table: alkali metals, alkaline earth metals, transition metals, metalloids, nonmetals, halogens, and noble gases.
- Certain groups of elements always form ions with the same charge. See page 74.
- Compounds containing ions conduct electricity if the ions can move.
- **Writing formulas for ionic compounds:** the charges have to balance out to neutral. Use the crossing method with the charges, but reduce if there's a common factor. For example:
 Mg^{2+} and N^{3-} for Mg_3N_2 but Mg^{2+} and O^{2-} form MgO

Chapter 4: Naming compounds.

- **Ionic compounds have a metal and at least one non-metal, and are stuck together by charges.**
- **Covalent compounds have only non-metals, and are held together by bonds.**
- Anions of a single element use the start of the element name, and end in -ide. Nitrogen with a negative charge becomes nitride. Oxygen with a negative charge is oxide.
- Ions can be a group of elements all stuck together by covalent bonds, that overall have a charge.
- If the group has an element attached to some number of oxygens, then it ends in -ate, and starts with the beginning of the name of the non-oxygen element. For example, NO_3^- is nitrate, PO_4^{3-} is phosphate.
- You can have a different number of oxygens, which have different naming schemes. The one that ends in -ate is the most common one.

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John Hornsby, Terry McGinnis, Margaret Lial

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Beginning and Intermediate Algebra Margaret L. Lial,E. John Hornsby,2003 [Beginning and Intermediate Algebra](#)
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