



EARTH SCIENCE

INTRODUCTION

COMPONENTS OF STUDY

- Geology: Earth physical & historical
- Metgeology: Weather & Climate
- Oceanography: The oceans
- Astronomy: What is external to the atmosphere

THE FOUR SPHERES (See The Water Cycle)

- Lithosphere: The solid earth
- Hydrosphere: The liquid earth
- Atmosphere: The gaseous earth
- Biosphere: Where life exists

CONCERN IN OUR PHYSICAL ENVIRONMENT

- Resources: Water, soil, metallic & non-metallic elements, sunlight, sunlight, wind, etc.
- The capita use of metallic and non-metallic elements is 11 tons per year
- Renewable vs. non-renewable resources
- Environmental and population problems

MINERALS

- A mineral is a naturally occurring, inorganic, solid material with a definite atomic structure.
- An ore is a useful mineral resource.

A. Atom and Minerals

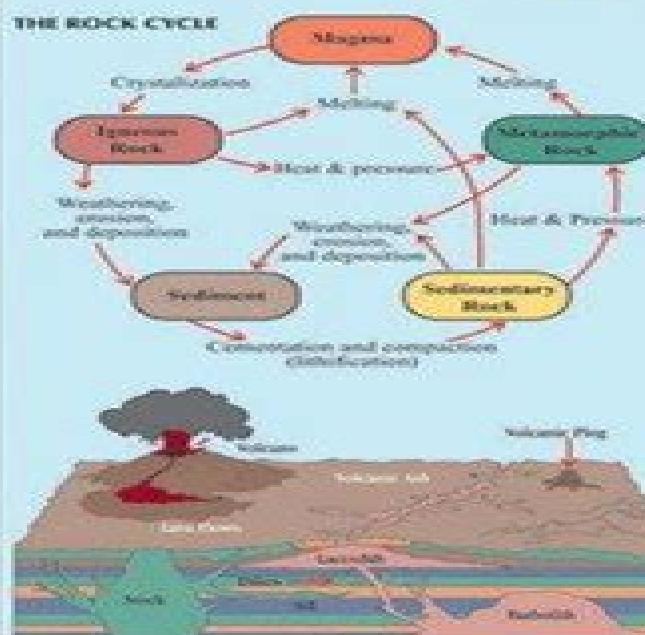
- Atoms in ancient Greek means "un-cut" i.e., the Greeks believed the atom to be the smallest particle of matter.
- Atoms are composed of neutrons, protons and electrons.
- Atomic number: Number of protons in nucleus.
- Atomic weight: Average weight of an atom.
- Isotopes: Atoms with the identical atomic number, but different numbers of neutrons in the nucleus.
- Forces: Strong, weak, electromagnetic and gravity.
- Properties of Minerals**
 - Color: Appearance or quality of light from surface.
 - Texture: Nature of light and cause of color.
 - Streak: Color of material in powdered form.
 - HARDNESS: Mohs' Scale (See Figure Below).
 - Scratch Test: Higher numbered materials can scratch lower numbered materials.
 - Shape is determined by cleavage, crystal form and structure.
- Specific Gravity: The ratio of the weight of a mineral to the weight of an equal volume of water.
- Density of water = 1 gram/cubic centimeter.
1 g/cm³ = 7.7, aluminum = 2.7, iron = 7.2.

MOM'S SCALE		MINERALS
10	Diamond	
9	Corundum	
8	Topaz	
7	Quartz	
6	feldspar	
5	Spatite	
4	Fluorite	
3	Garnet	
2	Pyrope	
1	Redite	
0	Table	

MINERAL CONCEPTS

- Silicates: Minerals with silicon and oxygen.
- Silicates are a large class which forms a pyramidal-shaped structure with oxygen.
- Other forms include simple non-silicate minerals, oxides, sulfides, noble elements, 2-3 element hydrocarbons and quartz.
- Non-Silicates
- Make up one-fourth of continental crust.
- Carbides: Minerals with carbon and oxygen including calcium, those which are present limestone (calcium) and marble (calcium carbonate).
- Oxides: Oxygen-banded solids, i.e., iron sulfides (FeS), sulfates (MgSO₄), halides (CaCl₂, F).
- Halite: Mineral form of salt.
- Corundum: Plasma, calcium.
- Noble metals: Iron, zinc, gold, silver, nickel.

THE ROCK CYCLE



ROCK TYPES

METAMORPHIC ROCKS

A. Formation

- Metamorphism**: To change form.
- Metamorphic rock**: Material under intense pressure, high temperatures.
- Contact metamorphism**: Metamorphism caused by proximity to magma.
- Low-grade metamorphism**: Shallow metamorphism.
- High-grade metamorphism**: Deeper metamorphism.
- foliation**: During metamorphism, parallel mineral lines are formed than metamorphic rocks.
- Metamorphic structures**: Features of material that change when heat and pressure.
- foliation**: Features of material that change when heat and pressure.

B. Agents of Metamorphism

- Heat**
 - most important agent
 - possible energy for chemical reactions
 - For example, clay converts into a mineral at great temperature
- Pressure**
- Chemical Activity**: Most common chemical agent in nature

C. Types of Metamorphic Rocks

- Metamorphic**: Fine grained, dark, shiny rock with randomly arranged minerals.
- Schist**: Fine grained, often grey, foliated rock with many strong cleavage planes of mica minerals formed by pressure.
- Migmatite**: Granular or mica-schist rock.
- Phyllite**: Shiny, foliated rock more strongly foliated than schist.
- Slate**: Foliated rock, more strongly foliated and w/ higher metamorphic grade than phyllite.
- Gneiss**: Foliated, banded rock w/ the highest metamorphic grade.
- Quartzite**: Very hard, granular quartz rock, formed from sandstone.

ROCKS' REACTION SERIES

It is a mineral, which has already formed, reacts w/ other minerals, it will react w/ the remaining magma to produce the next mineral in the sequence. For example, olivine reacts w/ plagioclase, plagioclase reacts w/ remaining magma to form pyroxene.

D. High-Mg Shale and High-Silicate Shale

- Calcareous rocks w/ other minerals form small crystals w/ foliation.
- These rocks w/ sodium w/ the reaction to become more and more sodium-rich.

E. End of Cooling

- Without everything is altered cool, remaining magma w/ have high olivine content, and quartz w/ foliation.
- Without cooling w/ temperature, everything that reacted at the same time will remain by the place for one another (plagioclase and quartz close near each other for mid-granite).



Earth Science High School Essential Question

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