

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

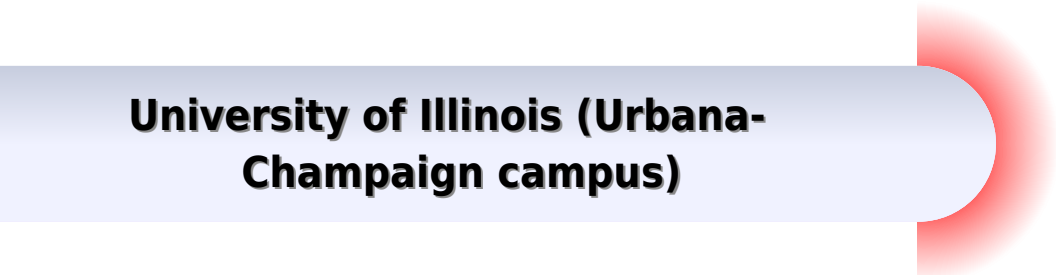
Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
 $V_1 = 3.8 \text{ L}$; $P_1 = 2.71 \text{ atm}$
 $V_2 = 1.47 \text{ L}$; $P_2 = ? \text{ Atm}$
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1 V_1 = P_2 V_2$
Isolate for P_2 : $P_2 = (P_1 V_1)/V_2$
 $P_2 = (2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L}) = 7.0 \text{ atm}$ (L. cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
 $V_1 = 1.49 \text{ L}$, $T_1 = 34.75 \text{ degrees Celsius}$
 $V_2 = ? \text{ L}$, $T_2 = 78.41 \text{ degrees Celsius}$
We need to convert from Celsius to Kelvins before proceeding.
 $T_1 = 34.75 + 273 = 307.75 \text{ K}$
 $T_2 = 78.41 + 273 = 351.41 \text{ K}$
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$
Get rid of P
 $V_1/T_1 = V_2/T_2$
Isolate V_2 : $V_2 = (V_1/T_1)T_2$
 $V_2 = (1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K}) = 1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
 $PV = nRT$ will work. (n = moles and R = 0.08206 L atm/K mol)
 $P = 1.04 \text{ atm}$, $V = ? \text{ L}$, n = OOPS we need to convert from grams to moles!, $T = 84.7 \text{ degrees Celsius} + 273 = 357.7 \text{ K}$
 n : Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$
Isolate V : $V = (nRT)/P = (4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

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Champaign campus)**



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shown in figure 17 2 2 17 2 2 and introduced in the text figure 17 2 2 17 2 2 this simplified drawing of the cardiovascular

system shows its main structures the heart is shown in the chest in red

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web sep 12 2023 the circulatory system is divided into two separate loops the shorter pulmonary circuit that exchanges blood between the heart and the lungs for oxygenation and the longer systemic circuit that distributes blood throughout all other systems and tissues of the body

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web sep 5 2023 circulatory system system that transports nutrients respiratory gases and metabolic products throughout a living organism circulation includes the intake of metabolic materials the movement of these materials to and from tissues and organs and the return of harmful by products to the environment