

# AP CHEMISTRY

Volume 1: Your Second Year Study

## Properties of Solids

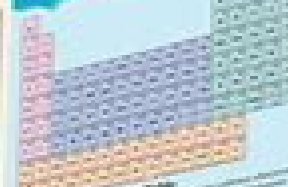
- **Crystalline Solids:** Atoms, ions, or molecules are arranged in a regular, repeating three-dimensional pattern. They have sharp melting points and distinct cleavage planes.
- **Amorphous Solids:** Atoms, ions, or molecules are arranged in a disordered, non-repeating pattern. They do not have sharp melting points and do not cleave along specific planes.
- **Intermolecular Forces:** The strength of these forces determines the melting point and hardness of a solid.
  - **London Dispersion Forces:** Present in all solids, they are the weakest intermolecular forces.
  - **Dipole-Dipole Interactions:** Present in polar solids, they are stronger than London forces.
  - **Hydrogen Bonding:** Present in solids containing N-H, O-H, or F-H bonds, they are the strongest intermolecular forces.
- **Unit Cells:** The smallest repeating unit of a crystal lattice. Common types include simple cubic, body-centered cubic (BCC), and face-centered cubic (FCC).
- **Packing Efficiency:** The fraction of space occupied by atoms in a unit cell. FCC and HCP have the highest packing efficiency (74%), while BCC has a lower efficiency (68%).

## Types of Chemical Calculations

Key to Answer on the Exam

- **Stoichiometry:** Calculations involving the relative amounts and masses of substances in a chemical reaction.
- **Thermochemistry:** Calculations involving heat energy changes during chemical reactions.
- **Equilibrium:** Calculations involving the concentrations of reactants and products at equilibrium.
- **Acid-Base:** Calculations involving pH, pOH, and the concentrations of acids and bases.
- **Solubility:** Calculations involving the concentrations of ions in saturated solutions.
- **Electrochemistry:** Calculations involving the relationship between electrical energy and chemical reactions.

## Periodic Table & Trends

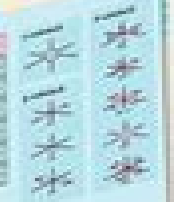


### Properties of the Periodic Table

- **Atomic Number:** The number of protons in the nucleus of an atom, denoted by Z.
- **Mass Number:** The sum of protons and neutrons in the nucleus, denoted by A.
- **Isotopes:** Atoms of the same element with different numbers of neutrons.
- **Periodicity:** The repeating nature of chemical and physical properties of elements as a function of their atomic number.

## Electronegativity Table

- **Electronegativity:** A measure of an atom's ability to attract and hold onto electrons in a chemical bond.
- **Trends:** Electronegativity increases from left to right across a period and from bottom to top within a group.



- **Atomic Size:** The distance from the nucleus to the outermost shell of electrons.
- **Ionic Size:** The distance from the nucleus to the outermost shell of electrons in an ion. Cations are smaller than their parent atoms, while anions are larger.
- **Molecular Size:** The distance between the nuclei of two atoms in a molecule.

## Thermodynamics

**First Law of Thermodynamics:** Energy cannot be created or destroyed, only converted from one form to another.  $\Delta E = q + w$

**Second Law of Thermodynamics:** The total entropy of an isolated system can never decrease over time.  $\Delta S_{univ} \geq 0$

**Gibbs Free Energy:** A measure of the maximum work that can be obtained from a thermodynamic process at constant temperature and pressure.  $\Delta G = \Delta H - T\Delta S$

## Equilibrium & Solubility

**Chemical Equilibrium:** A state in which the concentrations of reactants and products remain constant over time.  $K_{eq} = \frac{[products]}{[reactants]}$

**Le Chatelier's Principle:** If a system at equilibrium is subjected to a change in concentration, temperature, or pressure, the system will adjust to counteract the change.

Compound	K <sub>sp</sub>
AgCl	1.8 × 10 <sup>-10</sup>
PbCl <sub>2</sub>	1.7 × 10 <sup>-5</sup>
CaF <sub>2</sub>	3.9 × 10 <sup>-11</sup>
Fe(OH) <sub>3</sub>	4 × 10 <sup>-38</sup>
Al(OH) <sub>3</sub>	3 × 10 <sup>-34</sup>
Cu(OH) <sub>2</sub>	2.2 × 10 <sup>-20</sup>
Mg(OH) <sub>2</sub>	1.2 × 10 <sup>-17</sup>
Sn(OH) <sub>4</sub>	1 × 10 <sup>-46</sup>
Bi(OH) <sub>3</sub>	4 × 10 <sup>-40</sup>
Cr(OH) <sub>3</sub>	6.3 × 10 <sup>-31</sup>
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# Ap Chemistry Quick Study Academic Wallchart Common

**WJ Hussar**



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