

## **AP Biology Study Guide Unit 1**

### **Properties of Water:**

1. Polarity - allows H<sub>2</sub>O to form Hydrogen Bonds
2. Cohesion - Water H bonds to itself, creating surface tension
3. Adhesion - Water adheres to other polar substances
  - a. Allows easy transport of water against Gravity
4. Universal Solvent - Water can dissolve Polar and Ionic Compounds
  - a. This allows for reactions between Aqueous Solutions
5. Expands when Frozen - this means that Ice floats
6. High Specific Heat - water can absorb or release a large amount of heat with only a slight change in its own temperature (caused by cohesion)
  - a. High Heat of Vaporization - only warmest molecules can evaporate, so this causes evaporative cooling
7. Can be regulated by non-polar hydrophobic substances
8. Neutral PH → can form acids or bases if substances are added

Polarity - Difference in electronegativity between atoms bonded together create slightly charged ends of the molecules they form

Hydrophobic Substance - Nonpolar substance that repels water

Hydrophilic Substance - Polar substance that attracts water molecules

### **CHON - Carbon, Hydrogen, Nitrogen, Oxygen**

- Makeup 99% of living matter
- Strong Biological Cornerstones Because:
  - Smallest Diameter/Valency for atoms of 1, 2, 3, 4 valence electrons
  - Form Stable Covalent Bonds
  - Can form double and triple bonds
  - Versatile

### **Biological Transition Metals - Mn, Fe, Co, Cu, Zn**

- Can exist in multiple oxidation states → can be both oxidized and reduced

### **Biological Ions: Na<sup>+</sup>, Mg<sup>2+</sup>, Cl<sup>-</sup>, K<sup>+</sup>, Ca<sup>2+</sup>**

- Used to maintain homeostasis and osmotic concentration

Phosphorous/Sulphur - Energy Carriers which form strong high energy bonds

### **5 Foundation Biological Molecules:**

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**Mark Anestis, Kellie Ploeger Cox**



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