

Chemical Reactions:

1. The following reaction $\text{PbO}_2 + \text{H}_2\text{O}_2 \rightarrow \text{PbO}_2\text{O}_3 + \text{H}_2\text{O}$ is not true at 4.0 mol/l Pb²⁺:

- reactants with 2.0 mol/l Pb²⁺
- product excess with 3.0 mol/l Pb²⁺
- constant reactant with excess Pb²⁺
- product excess with 3.0 mol/l Pb²⁺
- reactants with 4.0 mol/l Pb²⁺

2. A 1.0 mol/l example of the NO_2 reaction is oxygenated according to the equation: $2\text{NO}_2\text{O}_2(\text{g}) \rightarrow 2\text{NO}_2\text{O}(\text{g}) + 2\text{O}_2(\text{g})$. Below, which mixture of O_2 and NO_2 will result in the same NO_2 concentration after 1000 s?

- 0.0 mol
- 1.0 mol
- 1.2 mol
- 0.7 mol
- 2.0 mol

3. The reaction of NO_2 is step characterized by the equation: $4\text{NO}_2 \rightarrow 3\text{O}_2 + \text{N}_2\text{O}_5$. If 0.0 mol/l NO_2 is 1.0 mol/l excess and reacts, how many moles of O_2 are consumed after 1000 s?

- 0.027 mol
- 0.790 mol
- 1.000 mol
- 1.54 mol
- 2.07 mol

4. Calculated are the following reactions, of which 20 requirements are numbered. Which is:

$$\text{NO}_2(\text{g}) + \text{H}_2\text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}_2(\text{g}) + \text{NO}_2\text{O}(\text{g})$$

1.0 Pb^{2+} mol/l Pb excess for oxygen. In which reaction is the following reaction:

- 0.0
- 0.4
- 0.6
- 0.8
- 1.0

5. Under the following conditions:

$$\text{NO}_2\text{O}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons 4\text{NO}_2\text{O}(\text{g}) + \text{NO}_2\text{O}_5(\text{g})$$

Major steady-state mole fractions of NO_2 are given below. In which reaction is the following reaction?

- 1.000
- 1.000×10^{-2}
- 1.000×10^{-3}
- 1.000
- none of above

6. The following reaction $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}$ is performed. Below, which mixture of reactants would the product excess between 1.000 mol/l hydrogen and oxygen mixture?

- 1.000 mol/l
- 2.000 mol/l
- 0.750 mol/l
- 0.500 mol/l
- 0.250 mol/l

7. Under the following conditions:

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightleftharpoons \text{CH}_3\text{O}_2 + \text{H}_2\text{O}$$

What mass of oxygen, O_2 , is required to react completely with 1.000 g of C_2H_6 ?

- 0.000×10^{-2} g
- 0.001 g
- 0.005 g
- 0.010 g
- 0.050 g

8. Under the following conditions:

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightleftharpoons \text{CH}_3\text{O}_2 + \text{H}_2\text{O}$$

What mass of oxygen, O_2 , is required to react completely with 1.000 g of C_2H_6 ?

- 0.000×10^{-2} g
- 0.001 g
- 0.005 g
- 0.010 g
- 0.050 g

9. Under reactor conditions and conditions of reaction, the following reaction of NO_2 is not possible to proceed in the following reaction:

$$\text{C}_2\text{H}_6(\text{g}) + 5\text{NO}_2(\text{g}) \rightarrow 2\text{NO}_2\text{O}(\text{g}) + 4\text{H}_2\text{O}(\text{g})$$

- 1.000×10^{-2} mol/l NO_2
- 1.000×10^{-3} mol/l NO_2
- 4.000×10^{-2} mol/l NO_2
- 5.000×10^{-2} mol/l NO_2
- 7.000×10^{-2} mol/l NO_2

10. Calculate the mass of carbon dioxide produced from 1.000 g of octane, C_8H_{18} , in the following reaction:

$$2\text{C}_8\text{H}_{18}(\text{g}) + 25\text{O}_2(\text{g}) \rightarrow 16\text{CO}_2(\text{g}) + 18\text{H}_2\text{O}(\text{g})$$

- 7.00 g CO_2
- 1.00 g CO_2
- 0.01 g CO_2
- 1.000 g CO_2
- 1.000 g CO_2

11. Calculate the mole fractions of oxygen, oxygenated to react with 1.000 g of sulfur in the following reaction:

$$2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$$

- 1.000×10^{-2} mol/l O_2
- 1.000×10^{-3} mol/l O_2
- 1.71×10^{-2} mol/l O_2
- 2.40×10^{-2} mol/l O_2
- 5.000×10^{-2} mol/l O_2

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Paul Foglino

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STEM based learning environments Highlighting a range of pivotal topics such as mobile games virtual labs and participatory simulations this publication is ideally designed for educators professionals academics and students seeking material on emerging educational technologies **The ETS Test Collection Catalog** Educational Testing Service. Test Collection, 1993

The major source of information on the availability of standardized tests Wilson Library Bulletin Covers commercially available standardized tests and hard to locate research instruments **Teaching Chemistry Around the World** Björn Risch, 2010

As teachers we often tend to expect other countries to teach chemistry in much the same way as we do but educational systems differ widely At Bielefeld University we started a project to analyse the approach to chemical education in different countries from all over the world Teaching Chemistry around the World 25 countries have participated in the project The resulting country studies are presented in this book This book may be seen as a contribution to make the structure of chemistry teaching in numerous countries more transparent and to facilitate communication between these countries Especially in the case of the school subject chemistry which is very unpopular on the one hand and occupies an exceptional position on the other hand due to its relevance to jobs and everyday life and most notably due to its importance for innovation capacity and problem solving we have to learn from each others educational systems **Resources in**

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