

## Chemistry 2<sup>nd</sup> Semester Final Review **KEY**

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

*Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)*

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)  
 $V_1 = 3.8 \text{ L}$ ;  $P_1 = 2.71 \text{ atm}$   
 $V_2 = 1.47 \text{ L}$ ;  $P_2 = ? \text{ Atm}$   
Since all units are standard (L, atm), we do not need to convert the units before proceeding.  
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$   
T goes away (not mentioned in problem)  
 $P_1 V_1 = P_2 V_2$   
Isolate for  $P_2$ :  $P_2 = (P_1 V_1)/V_2$   
 $P_2 = (2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L}) = 7.0 \text{ atm}$  (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)  
 $V_1 = 1.49 \text{ L}$ ,  $T_1 = 34.75 \text{ degrees Celsius}$   
 $V_2 = ? \text{ L}$ ,  $T_2 = 78.41 \text{ degrees Celsius}$   
We need to convert from Celsius to Kelvins before proceeding.  
 $T_1 = 34.75 + 273 = 307.75 \text{ K}$   
 $T_2 = 78.41 + 273 = 351.41 \text{ K}$   
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$   
Get rid of P  
 $V_1/T_1 = V_2/T_2$   
Isolate  $V_2$ :  $V_2 = (V_1/T_1)T_2$   
 $V_2 = (1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K}) = 1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?  
The combined gas law won't work here. You only have one condition and nothing changes.  
 $PV = nRT$  will work. ( $n$  = moles and  $R$  = 0.08206 L atm/K mol)  
 $P = 1.04 \text{ atm}$ ,  $V = ? \text{ L}$ ,  $n$  = OOPS we need to convert from grams to moles!,  $T = 84.7 \text{ degrees Celsius} + 273 = 357.7 \text{ K}$   
 $n$ : Hydrogen gas is diatomic!  
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$   
Isolate  $V$ :  $V = (nRT)/P = (4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

# 2nd Semester Chemistry Final Exam Review Answers

**Harold Goldwhite, John R. Spielman**



## 2nd Semester Chemistry Final Exam Review Answers:

*Artificial Intelligence in Education. Posters and Late Breaking Results, Workshops and Tutorials, Industry and Innovation Tracks, Practitioners, Doctoral Consortium and Blue Sky* Andrew M. Olney, Irene-Angelica Chounta, Zitao Liu, Olga C. Santos, Ig Ibert Bittencourt, 2024-07-01 This volume constitutes poster papers and late breaking results presented during the 25th International Conference on Artificial Intelligence in Education AIED 2024 which took place in Recife Brazil during July 8 12 2024 The 18 full papers and 92 short papers were carefully reviewed and selected from 200 submissions They are organized in topical sections as follows Part One Blue Sky Industry Innovation and Practitioner Wide AIED and Late Breaking Results Part Two Late Breaking Results Doctoral Consortium Workshops and Tutorials Nursing School Entrance Exam Review William H. Brown, Frederick A. Bettelheim, Jerry March, 2004-04

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