

CHEMISTRY LAB: BEAN BAG ISOTOPES

Introduction

Studies of radioactivity at the beginning of the 20th century made it possible to investigate the actual structure and mass of atoms. Gradually, evidence began to build that atoms of the same element could have different masses. These atoms were called isotopes. How are isotopes distinguished from one another? What is the average atomic mass of an element that has different isotopes?

Background

Isotopes are defined as atoms that have the same number of protons but different numbers of neutrons. Since the identity of an element depends only on the number of protons (atomic number), isotopes have the same chemical properties. Isotopes are thus chemically identical – they form the same compounds, undergo the same reactions, etc. Isotopes are distinguished from one another based on their mass number, defined as the sum of the number of protons and neutrons in the nucleus of the atoms.

Chlorine, for example, occurs naturally in the form of two isotopes, chlorine-35 and chlorine-37, where 35 and 37 represent the mass numbers of the isotopes. Each isotope of chlorine has a characteristic percent abundance in nature. Thus, whether it is analyzed from underground salt deposits or from seawater, the element chlorine always contains 75.8% chlorine-35 atoms and 24.2% chlorine-37 atoms. The atomic mass of an element represents a weighted average of the masses of the isotopes in a naturally occurring sample.

Purpose

The purpose of this activity is to investigate the mass properties and relative abundance of isotopes for the "bean bag" element (symbol Bg) and to calculate the atomic mass of this element.

Procedure

1. Cover your work area with paper towels. You will use the beans and want to keep them as clean as possible.
2. Sort the atoms in the "bean bag" element sample into three isotope groups (1, 2, and 3) according to the type of bean. Assume that each type of bean represents a different isotope and that each bean represents a separate atom.
3. Count the number of Bg atoms in each isotope group and record this number in your data table.
4. Measure the total mass of Bg atoms belonging to each isotope group. Record each mass in your data table. *Remember to use a weighing dish when measuring your atoms.

Calculations – remember to show all work

1. Determine the average mass of each Bg isotope and record this value in your data table.
2. Determine the percent abundance of each Bg isotope and record this value in your data table.
3. Calculate the atomic mass for the bean bag element.

Bean Bag Isotopes Lab Answers

Marcel A. Müller

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