

## Chemistry 2<sup>nd</sup> Semester Final Review KEY

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)

$$V_1 = 3.8 \text{ L}; P_1 = 2.71 \text{ atm}$$

$$V_2 = 1.47 \text{ L}; P_2 = ? \text{ atm}$$

Since all units are standard (L, atm), we do not need to convert the units before proceeding.

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

T goes away (not mentioned in problem)

$$P_1 V_1 = P_2 V_2$$

Isolate for P2:  $P_2 = (P_1 V_1) / V_2$

$$P_2 = (2.71 \text{ atm} \times 3.8 \text{ L}) / (1.47 \text{ L}) = 7.0 \text{ atm} \text{ (L cancelled and it's 2 sig figs)}$$

2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)

$$V_1 = 1.49 \text{ L}, T_1 = 34.75 \text{ degrees Celsius}$$

$$V_2 = ? \text{ L}, T_2 = 78.41 \text{ degrees Celsius}$$

We need to convert from Celsius to Kelvins before proceeding.

$$T_1 = 34.75 + 273 = 307.75 \text{ K}$$

$$T_2 = 78.41 + 273 = 351.41 \text{ K}$$

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

Get rid of P

$$V_1 / T_1 = V_2 / T_2$$

Isolate V2:  $V_2 = (V_1 / T_1) T_2$

$$V_2 = (1.49 \text{ L} / 307.75 \text{ K}) (351.41 \text{ K}) = 1.70 \text{ L}$$

3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?

The combined gas law won't work here. You only have one condition and nothing changes.

$PV = nRT$  will work. ( $n$ = moles and  $R = 0.08026 \text{ L atm/K mol}$ )

$P = 1.04 \text{ atm}$ ,  $V = ? \text{ L}$ ,  $n =$  OOPS we need to convert from grams to moles!,  $T = 84.7 \text{ degrees Celsius} + 273 = 357.7 \text{ K}$

$n$ : Hydrogen gas is diatomic!

$$8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$$

$$\text{Isolate V: } V = (nRT) / P = (4.193069307 \text{ mol H}_2 \times 0.08026 \text{ L atm/K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$$

4. What volume is occupied by 56.75 g of oxygen gas at STP?

# Nd Semester Final Review Guide Chemistry

**Lingsheng Yao**

## Nd Semester Final Review Guide Chemistry:

Teaching First Year Chemistry Joseph Otto Frank,1927    **Report** International Clearinghouse on Science and Mathematics Curricular Developments,1972    Resources in Education ,1978    **Year-book** University of Southern California,1903    A Guide to Educational Measurements Harlan Cameron Hines,1923    **Paperbound Books in Print** ,1967    *Catalog of Carleton College for the Academic Year ...* Carleton College (Northfield, Minn.),1909    **Journal of Chemical Education** ,1925 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers    **Host Bibliographic Record for Boundwith Item Barcode 30112062967754 and Others** ,1903    *Medical Education, Medical Colleges and the Regulation of the Practice of Medicine in the United States and Canada* ,1891    Circular of Information University of Southern California,1904    Medical Education, Medical Colleges and the Regulation of the Practice of Medicine in the United States and Canada, 1765-1891 Illinois State Board of Health,John Henry Rauch,1891    **Annual Catalogue of Elizabethtown College** Elizabethtown College,1928    **Research in Education** ,1969-05    **School Science and Mathematics** ,1928    **The National Guide to Educational Credit for Training Programs** American Council on Education,2005 Highlights over 6 000 educational programs offered by business labor unions schools training suppliers professional and voluntary associations and government agencies    **Ohio University Bulletin** Ohio University,1913    University of Illinois Bulletin ,1910    Catalogue of Bucknell University Bucknell University,1930    **Peterson's Annual Guides to Graduate Study** ,1983

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### **Table of Contents Nd Semester Final Review Guide Chemistry**

1. Understanding the eBook Nd Semester Final Review Guide Chemistry
  - The Rise of Digital Reading Nd Semester Final Review Guide Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Nd Semester Final Review Guide Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Nd Semester Final Review Guide Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Nd Semester Final Review Guide Chemistry
  - Personalized Recommendations
  - Nd Semester Final Review Guide Chemistry User Reviews and Ratings
  - Nd Semester Final Review Guide Chemistry and Bestseller Lists

5. Accessing Nd Semester Final Review Guide Chemistry Free and Paid eBooks
  - Nd Semester Final Review Guide Chemistry Public Domain eBooks
  - Nd Semester Final Review Guide Chemistry eBook Subscription Services
  - Nd Semester Final Review Guide Chemistry Budget-Friendly Options
6. Navigating Nd Semester Final Review Guide Chemistry eBook Formats
  - ePUB, PDF, MOBI, and More
  - Nd Semester Final Review Guide Chemistry Compatibility with Devices
  - Nd Semester Final Review Guide Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Nd Semester Final Review Guide Chemistry
  - Highlighting and Note-Taking Nd Semester Final Review Guide Chemistry
  - Interactive Elements Nd Semester Final Review Guide Chemistry
8. Staying Engaged with Nd Semester Final Review Guide Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Nd Semester Final Review Guide Chemistry
9. Balancing eBooks and Physical Books Nd Semester Final Review Guide Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Nd Semester Final Review Guide Chemistry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Nd Semester Final Review Guide Chemistry
  - Setting Reading Goals Nd Semester Final Review Guide Chemistry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nd Semester Final Review Guide Chemistry
  - Fact-Checking eBook Content of Nd Semester Final Review Guide Chemistry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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