

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
 $V_1 = 3.8 \text{ L}$; $P_1 = 2.71 \text{ atm}$
 $V_2 = 1.47 \text{ L}$; $P_2 = ? \text{ Atm}$
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1 V_1 = P_2 V_2$
Isolate for P_2 : $P_2 = (P_1 V_1)/V_2$
 $P_2 = (2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L}) = 7.0 \text{ atm}$ (L cancelled and it's 2 sig figs)
2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
 $V_1 = 1.49 \text{ L}$, $T_1 = 34.75 \text{ degrees Celsius}$
 $V_2 = ? \text{ L}$, $T_2 = 78.41 \text{ degrees Celsius}$
We need to convert from Celsius to Kelvins before proceeding.
 $T_1 = 34.75 + 273 = 307.75 \text{ K}$
 $T_2 = 78.41 + 273 = 351.41 \text{ K}$
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$
Get rid of P
 $V_1/T_1 = V_2/T_2$
Isolate V_2 : $V_2 = (V_1/T_1)T_2$
 $V_2 = (1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K}) = 1.70 \text{ L}$
3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
 $PV = nRT$ will work. (n = moles and $R = 0.08206 \text{ L atm/K mol}$)
 $P = 1.04 \text{ atm}$, $V = ? \text{ L}$, $n = \text{OOPS we need to convert from grams to moles!}$, $T = 84.7 \text{ degrees Celsius} + 273 = 357.7 \text{ K}$
 n : Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$
Isolate V : $V = (nRT)/P = (4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$
4. What volume is occupied by 56.75 g of oxygen gas at STP?

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Lingsheng Yao



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Teaching First Year Chemistry Joseph Otto Frank, 1927 **Report** International Clearinghouse on Science and Mathematics Curricular Developments, 1972 Resources in Education , 1978 **Year-book** University of Southern California, 1903 A Guide to Educational Measurements Harlan Cameron Hines, 1923 **Paperbound Books in Print** , 1967 *Catalog of Carleton College for the Academic Year ...* Carleton College (Northfield, Minn.), 1909 **Journal of Chemical Education** , 1925 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers **Host Bibliographic Record for Boundwith Item Barcode 30112062967754 and Others** , 1903 *Medical Education, Medical Colleges and the Regulation of the Practice of Medicine in the United States and Canada* , 1891 Circular of Information University of Southern California, 1904 Medical Education, Medical Colleges and the Regulation of the Practice of Medicine in the United States and Canada. 1765-1891 Illinois State Board of Health, John Henry Rauch, 1891 **Annual Catalogue of Elizabethtown College** Elizabethtown College, 1928 **Research in Education** , 1969-05 **School Science and Mathematics** , 1928 The National Guide to Educational Credit for Training Programs American Council on Education, 2005 Highlights over 6 000 educational programs offered by business labor unions schools training suppliers professional and voluntary associations and government agencies **Ohio University Bulletin** Ohio University, 1913 University of Illinois Bulletin , 1910 Catalogue of Bucknell University Bucknell University, 1930 **Peterson's Annual Guides to Graduate Study** , 1983

Nd Semester Final Review Guide Chemistry Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the energy of words has been much more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such may be the essence of the book **Nd Semester Final Review Guide Chemistry**, a literary masterpiece that delves deep into the significance of words and their impact on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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