

**1st Semester Exam**  
**Chemistry Multiple Choice Section**  
**Please DO NOT write on this form!!**

1. What type of scientist would you consider a researcher project which uses previous knowledge of virus structure to find a cure for AIDS?  
a. pure sciences    b. applied sciences    c. logical sciences    d. practical sciences    e. technological sciences
2. If you want to conduct an experiment which will measure the effect of temperature on the composition rate for water, which of the following is the **BEST** procedure for this experiment?  
a. measure 20, 40, and 60 mL of water into 3 beakers, heat for 10 minutes at 60°C, measure the new volume after 10 minutes  
b. measure 40 mL of water into 3 beakers, heat one for 5 minutes, use the 15 mL, use the other for 25 min, measure the amount of water remaining in each beaker  
c. measure 25 mL of water into 3 beakers, place the 3 beakers on 25°C, 50°C, and 75°C heating pads, heat all three beakers for 30 minutes, measure the amount of water remaining in each beaker
3. Express 541,800 in scientific notation.  
a. 541,800    b.  $5.41 \times 10^5$     c.  $5.41 \times 10^6$     d.  $5.40 \times 10^5$
4. Express  $3.09 \times 10^{-5}$  in decimal notation.  
a. 0.000309    b. 3050    c. 000003    d. 0.0005    e. 305
5. Mass gravity affects an object is known as its \_\_\_\_\_  
a. weight    b. mass    c. volume    d. density    e. energy
6. The two major characteristics of matter are that it  
a. has definite mass and shape    c. combines with other matter in compounds  
b. has mass and occupies space    d. is made up of atoms and molecules
7. A substance which can not be broken down by chemical means is known as \_\_\_\_\_  
a. element    b. compound    c. homogeneous mixture    d. heterogeneous mixture    e. solution
8. The fact that alkaline metals are able to react violently with acids is known as a  
a. physical change    b. physical property    c. chemical change    d. chemical property
9. When ice melts to become water, this is known as a  
a. physical change    b. physical property    c. chemical change    d. chemical property
10. What were the results and conclusions of Lavoisier's Candle Experiment?  
a. as the candle burned, mass was lost, therefore matter can be destroyed  
b. as the candle burned, mass was lost, therefore matter can be changed into energy  
c. as the candle burned, the react weighed the same as the original candle, therefore matter is conserved  
d. the flame contains matter
11. Which of the following is a good example of a homogeneous mixture?  
a. chocolate chip cookie    b. Kool-Aid    c. iron    d. water
12. Objects with the same charge  
a. attract each other    b. repel each other    c. have no effect on each other    d. not enough info
13. Which of the following is the best choice for measuring volume of a liquid?  
a. graduated cylinder    b. electronic balance    c. beaker    d. graduated cylinder    e. flask
14. What should you do **BEST** if you were to spray acid all over yourself?  
a. go to the nearest station, place your face above the sprayers, and turn it on with your eyes open  
b. stop, drop, and roll  
c. report the incident to your instructor  
d. run quickly under the shower head and pull the handle to turn it on
15. As the temperature of an object increases,  
a. the particles move faster    c. it loses heat  
b. the particles move more slowly    d. it releases cold waves
16. Metals are used to make pots and pans because they are good \_\_\_\_\_ of heat.  
a. generators    b. conductors    c. insulators    d. holders    e. absorbers
17. The best example of kinetic energy is the following list is:  
a. a wound up spring    c. an object sitting on a table  
b. an electron orbiting around a nucleus    d. an atom at absolute zero
18. Which of the following is absolute zero?  
a. 0 K    b. 0°C    c. 0°F    d. a,b, & c

# Nd Semester Exam Chemistry Form

**University of Michigan**

The logo of the University of Michigan, featuring a red circular emblem with a white 'U' and 'M' inside, positioned to the right of the text.

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