

Acids & Bases HOMEWORK

Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrogen fluoride**
- b. **HI**: **Hydrogen iodide**
- c. **HNO₃**: **nitric acid**
- d. **HClO₄**: **perchloric acid**
- e. **HBrO₃**: **hypobromous acid**
- f. **HClO₄**: **chloric acid**
- g. **H₂SO₄**: **hydrochloric acid**
- h. **H₂SO₄**: **hydrobromic acid**
- i. **H₂SO₄**: **hydroiodic acid**
- j. **H₂SO₄**: **hydrochloric acid**
- k. **H₂SO₄**: **hydrobromic acid**
- l. **H₂SO₄**: **hydroiodic acid**

2. Of the acids and bases listed below, write the ones that completely dissociate in water.

Acid	Base	Mixture	Neutral	Acidic	Basic	Neutral	Basic
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3. Write the dissociation for **H₂SO₄**.

4. **HClO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{H}_3\text{ClO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{ClO}_4 + \text{H}_3\text{O}^+$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating HClO_4 in a weak acid.

c. Dissolve **1.00\text{M}** NaClO_4 in a portion of water and a portion of **1.00\text{M}** **NaOH**. Dissolve **1.00\text{M}** HClO_4 in a portion of water.

d. In your **reaction** prediction, is HClO_4 acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because HClO_4 is a **strong** acid.

e. Identify the conjugate base of HClO_4 : **ClO_4^-**

f. Identify NaClO_4 as **strong**, **weak**, or **none**: **strong**. Because of common conjugate base ClO_4^- .

5. H^+ will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{H}_2\text{O} + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{O}^+ + \text{OH}^-$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating H_2O in a weak base.

c. Dissolve **1.00\text{M}** NaOH in a portion of water and a portion of **1.00\text{M}** **NaCl**. Dissolve **1.00\text{M}** H_2O in a portion of water.

d. In your **reaction** prediction, is H_2O acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because of a **strong** base.

e. Identify the conjugate acid of H_2O : **H_3O^+**

f. Identify NaOH as **strong**, **weak**, or **none**: **strong**.

6. **H₂SO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{SO}_4^+ + \text{H}_3\text{O}^+$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating H_2SO_4 in a weak acid.

c. Dissolve **1.00\text{M}** NaHSO_4 in a portion of water and a portion of **1.00\text{M}** **NaOH**. Dissolve **1.00\text{M}** H_2SO_4 in a portion of water.

d. In your **reaction** prediction, is H_2SO_4 acting as a **Brønsted-Lowry** acid or a **Brønsted-Lowry** base? **Brønsted-Lowry acid**. Because of a **strong** acid.

e. Identify the conjugate base of H_2SO_4 : **HSO_4^-**

f. Identify NaHSO_4 as **strong**, **weak**, or **none**: **strong**. Because of common conjugate base HSO_4^- in the **strong** base NaOH .

19 3 Acids And Bases Answer Key Chemistry

Ibrahim Sikder

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The Chemistry of the Several Natural and Artificial Heterogeneous Compounds Used in Manufacturing Porcelain, Glass, and Pottery Simeon Shaw, 1900 **Prentice Hall Chemistry**, 2000 **Chemistry Textbook for College and University**

USA Ibrahim Sikder, 2023-06-04 Chemistry Textbook USA Physical Chemistry for Engineering and Applied Sciences

Frank R. Foulkes, 2012-09-12 Physical Chemistry for Engineering and Applied Sciences is the product of over 30 years of teaching first year Physical Chemistry as part of the Faculty of Applied Science and Engineering at the University of Toronto. Designed to be as rigorous as compatible with a first year student's ability to understand the text presents detailed step by step derivations of the equations that permit the student to follow the underlying logic and of equal importance to appreciate any simplifying assumptions made or mathematical tricks employed. In addition to the 600 exercises and end of chapter problems the text is rich in worked non trivial examples many of which are designed to be inspiring and thought provoking. Step by step derivation of all equations enables the student to smoothly follow the derivation by sight and can be understood relatively easily by students with moderate skills and backgrounds in mathematics. Clear and accessible Physical Chemistry for Engineering and Applied Sciences includes The answers to all of the 112 worked examples 99 exercises following many of the worked examples and 496 end of chapter problems. Topics not normally seen in introductory physical chemistry textbooks ionic reaction rates activities and activity coefficients or not regularly explained in much detail electrochemistry chemical kinetics with an eye on industrial applications. Special appendices that provide detailed explanations of basic integration and natural logarithms for students lacking a background in integral calculus. An in depth chapter on electrochemistry in which activities and activity coefficients are used extensively as required for accurate calculations *A Dictionary of Chemistry and the Allied Branches of Other Sciences* Henry Watts, 1875 *General Chemistry* Ralph H. Petrucci, William S.

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Watts, 1905 **Chemistry** James Vincent Quagliano, L. M. Vallarino, 1969 **A Dictionary of applied chemistry v. 1,**

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