

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HBr~~ **hydrobromic acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO₃~~ **nitric acid**
- ~~H₂SO₄~~ **sulfuric acid**
- ~~Strontium hydroxide~~ **LiOH**
- ~~potassium hydroxide~~ **KOH**
- ~~calcium hydroxide~~ **CaOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **Li(OH)₄**

~~H₂O₂ for HYDROPEROXIDE acids it never will completely dissociate in water~~

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

~~HF~~

~~H₂SO₄~~

~~Mg(OH)₂~~

NaOH

NaHCO₃

~~NaOH~~

Na₂CO₃

3. What is the formula for hydronium? **H₃O⁺**

4. BaC_2O_4 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{BaC}_2\text{O}_4 + \text{H}_2\text{O} \rightleftharpoons \text{BaC}_2\text{O}_4^- + \text{HC}_2\text{O}_4^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because BaC_2O_4 is a weak acid.**
- Would you define BaC_2O_4 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is BaC_2O_4 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of BaC_2O_4 : **BaC_2O_4^-**
- Identify BaC_2O_4 as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's to give.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{F}^- + \text{H}_2\text{O} \rightleftharpoons \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither. **neither**

6. HSO_4^- will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{HSO}_4^- + \text{H}_2\text{O} \rightleftharpoons \text{SO}_4^{2-} + \text{H}_3\text{O}^+$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO_4^- is a weak acid.**
- Would you define HSO_4^- as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HSO_4^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO_4^- : **SO_4^{2-}**
- Identify HSO_4^- as polyprotic, amphoteric, or neither. **amphoteric, because it has a negative charge (and will take a proton) and it has H's to give (and will).**

19 3 Acids And Bases Answer Key Chemistry

Ibrahim Sikder



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