

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Identify ionization levels based on the pH of acids and bases and approximate concentrations.
2. Calculate hydrogen ion concentrations from pH and hydroxide ion concentrations from pOH values.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-------------------------------|---------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. basic solution | 6. neutral solution |
| | | 7. pH |

Key Equation

- | | |
|-------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

Part A Completion

Write the hydrogen ion concentration or hydroxide ion concentration of the solutions described. Round the concentration to the correct significant figures. Express values using scientific notation.

1. What substance is _____, contains hydrogen ions?
2. _____ is the concentration of those acids present in stomach acid equal to _____ molar.
3. _____ is the concentration of those acids present in _____ molar _____.
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19 2 Hydrogen Ions And Acidity Work Sheet Answers

Rod O'Connor



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