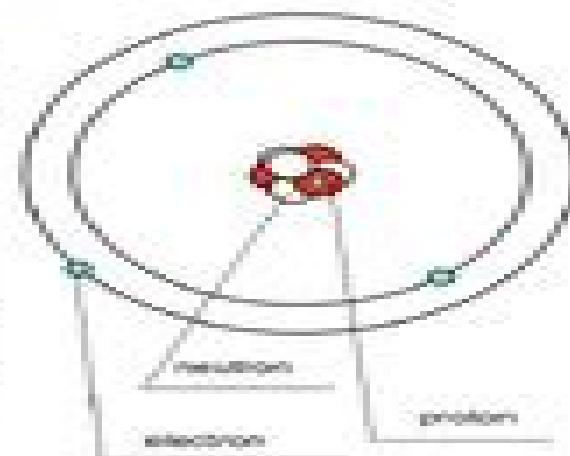


Answer Key

- PROTONS are found in the center of the atom in a region called the NUCLEUS. (slide 27)
- It is the number of protons that determine what ELEMENT an atom will be. (slide 23)
- The number of protons is called the ATOMIC NUMBER. (like an ID Number) (slide 29)
- The atomic number is a whole number found with the SYMBOL of the element. (slide 30)
- Protons have a mass of 1 amu and are one of the heavier pieces of an atom. (slide 1.1)
- Protons have a +1 charge. (slide 22)
- NEUTRONS are the other type of particle found in the center of the atom (nucleus). (slide 28)
- Neutrons have ZERO charge and have a mass of 1 amu. (slide 24-25)
- ELECTRONS are the smallest of the subatomic particles, found outside the nucleus. (slide 26)
- Electrons have virtually no significant MASS. (slide 26)
- Electrons have a -1 charge. (slide 27)
- Electrons can MOVE between atoms or be SHARED between atoms. (slide 27)
- It an atom is NEUTRAL, it will have the same number of electrons as protons. (slide 29)
- Electrons are involved in all CHEMICAL REACTIONS. (slide 25)

PARTICLE	MASS	CHARGE	LOCATION	ROLE
PROTON	1	+1	Nucleus	Identity of Element
NEUTRON	1	0	Nucleus	Hold Nucleus Together
ELECTRON	0	-1	Outside Nucleus	Reactivity



Atomic Structure Guided Answers

David Baud

Atomic Structure Guided Answers:

Atom Structure Guide Sophia Anderson, AI, 2025-02-22 Atom Structure Guide explores the core concepts of atomic physics providing a comprehensive look at the atom's structure and behavior through the lens of quantum mechanics and experimental findings. Consider that the quantum mechanical model replaced older classical notions with probabilities revolutionizing our understanding of chemical bonding and material properties. Further experimental techniques like spectroscopy and electron microscopy have offered crucial insights into the atom's internal structure. This book uniquely balances theoretical concepts like wave-particle duality and the Schrödinger equation with experimental evidence like spectroscopic measurements and microscopic images. It emphasizes how a solid grasp of atomic structure is essential for progress across scientific fields. The book starts with the historical development of atomic theory and systematically builds upon that foundation dedicating chapters to various spectroscopic techniques and concluding with the atom's role in chemical bonding, material properties, and quantum computing. The book's approach is designed to make complex topics accessible without sacrificing scientific accuracy, making it a valuable resource for students, researchers, and anyone interested in the fundamentals of physics and science.

Study Guide, Answer Book to Accompany Organic Chemistry Alan S. Wingrove, Robert L. Caret, 1981

Answer Guide for Chemistry, 2nd Edition Leonard W. Fine, 1978 **Study Guide with Answers to Selected Problems** H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview, activities, and practice problems to reinforce skills and a practice test. The answers section includes answers for all odd-numbered end-of-chapter exercises.

Microgravity News, 1999 *Readers' Guide to Periodical Literature*, 1915 *Platers' Guide*, 1919 *The Church Eclectic*, 1874 **Radiology**, 1926 **The Brass World and Platers**

Guide, 1922 **Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry** Iris Stovall, Donald J. DeCoste, 2000 *Chemistry, Study Guide* Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry, chemical equilibrium, and thermochemistry. Worked examples now number over 300 and exercises over 1460.

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