

1. Define the term 'isotope'. (1 mark)

The same element but different mass number.

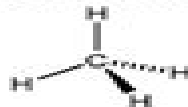
2. What type of covalent bond connects a hydrogen atom to an oxygen atom in a water molecule? (1 mark)

Polar

3. Define 'specific heat' of water. (1 mark)

The amount of heat absorbed or lost in order to alter the temperature of 1g of water by 1°C.

4. Draw a detailed picture of one molecule of methane. (2 marks)



One central carbon with covalent bonds to four atoms of hydrogen.

5. What is the pH of 0.01M NaOH? Show all working. (5 marks)

0.01M NaOH has an $[\text{OH}^-] = 10^{-2}$. (1 mark)

$[\text{H}^+] \times [\text{OH}^-] = 10^{-14}$ (1 mark)

Therefore, $\text{H}^+ = 10^{-12}$. (1 mark)

$\text{pH} = -\log_{10}[\text{H}^+]$ (1 mark)

therefore pH of 0.01M NaOH = $-(-12) = +12$ (1 mark)

6. Write the sequence of the mRNA molecule synthesized from a DNA template strand with the sequence 5'-ATCGTACCGTTA-3' (3 marks)

Answer: 5'-UAACGGUACGAU-3'

7. Determine the net charge of the predominant form of Glu at (a) pH 1.0, (b) pH 3.0, (c) pH 6.0 and (d) pH 11.0. The structure of Glu including pKa values for the ionisable groups is shown below. (4 marks)

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Merkhi



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Table of Contents Atmosphere End Of Module Test Answers

1. Understanding the eBook Atmosphere End Of Module Test Answers
 - The Rise of Digital Reading Atmosphere End Of Module Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Atmosphere End Of Module Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Atmosphere End Of Module Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Atmosphere End Of Module Test Answers
 - Personalized Recommendations
 - Atmosphere End Of Module Test Answers User Reviews and Ratings
 - Atmosphere End Of Module Test Answers and Bestseller Lists

5. Accessing Atmosphere End Of Module Test Answers Free and Paid eBooks
 - Atmosphere End Of Module Test Answers Public Domain eBooks
 - Atmosphere End Of Module Test Answers eBook Subscription Services
 - Atmosphere End Of Module Test Answers Budget-Friendly Options
6. Navigating Atmosphere End Of Module Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Atmosphere End Of Module Test Answers Compatibility with Devices
 - Atmosphere End Of Module Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Atmosphere End Of Module Test Answers
 - Highlighting and Note-Taking Atmosphere End Of Module Test Answers
 - Interactive Elements Atmosphere End Of Module Test Answers
8. Staying Engaged with Atmosphere End Of Module Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Atmosphere End Of Module Test Answers
9. Balancing eBooks and Physical Books Atmosphere End Of Module Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Atmosphere End Of Module Test Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Atmosphere End Of Module Test Answers
 - Setting Reading Goals Atmosphere End Of Module Test Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Atmosphere End Of Module Test Answers
 - Fact-Checking eBook Content of Atmosphere End Of Module Test Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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