

## 19.5

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                  |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain                | 6. acid rain     |
| 7. autoionization | 8. acid rain                | 9. acid rain     |

## Key Equation

- |                                    |                                    |
|------------------------------------|------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}^+]$ |

## Part A Completion

Write the equation for the autoionization of water. What is the  $\text{pH}$  of pure water at 25°C? What is the  $\text{pOH}$  of pure water at 25°C? What is the  $\text{pH}$  of a 0.1 M solution of HCl? What is the  $\text{pH}$  of a 0.1 M solution of NaOH? What is the  $\text{pH}$  of a 0.1 M solution of  $\text{H}_2\text{SO}_4$ ? What is the  $\text{pH}$  of a 0.1 M solution of  $\text{Ca(OH)}_2$ ?

What is the  $\text{pH}$  of a 0.1 M solution of HCl? What is the  $\text{pH}$  of a 0.1 M solution of NaOH? What is the  $\text{pH}$  of a 0.1 M solution of  $\text{H}_2\text{SO}_4$ ? What is the  $\text{pH}$  of a 0.1 M solution of  $\text{Ca(OH)}_2$ ?

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1.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2.  $\text{pH} = 7.0$
3.  $\text{pH} = 7.0$
4.  $\text{pH} = 1.0$
5.  $\text{pH} = 13.0$
6.  $\text{pH} = 1.0$
7.  $\text{pH} = 1.0$
8.  $\text{pH} = 1.0$
9.  $\text{pH} = 1.0$
10.  $\text{pH} = 1.0$
11.  $\text{pH} = 1.0$
12.  $\text{pH} = 1.0$
13.  $\text{pH} = 1.0$
14.  $\text{pH} = 1.0$
15.  $\text{pH} = 1.0$

# 19 2 Hydrogen Ions And Acidity Work Sheet Answers

**Advisory Council on College Chemistry**

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