

## Chemistry 2<sup>nd</sup> Semester Final Review KEY

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)

$$V_1 = 3.8 \text{ L}; P_1 = 2.71 \text{ atm}$$

$$V_2 = 1.47 \text{ L}; P_2 = ? \text{ atm}$$

Since all units are standard (L, atm), we do not need to convert the units before proceeding.

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

T goes away (not mentioned in problem)

$$P_1 V_1 = P_2 V_2$$

Isolate for P2:  $P_2 = (P_1 V_1) / V_2$

$$P_2 = (2.71 \text{ atm} \times 3.8 \text{ L}) / (1.47 \text{ L}) = 7.0 \text{ atm} \text{ (L cancelled and it's 2 sig figs)}$$

2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)

$$V_1 = 1.49 \text{ L}, T_1 = 34.75 \text{ degrees Celsius}$$

$$V_2 = ? \text{ L}, T_2 = 78.41 \text{ degrees Celsius}$$

We need to convert from Celsius to Kelvins before proceeding.

$$T_1 = 34.75 + 273 = 307.75 \text{ K}$$

$$T_2 = 78.41 + 273 = 351.41 \text{ K}$$

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

Get rid of P

$$V_1 / T_1 = V_2 / T_2$$

Isolate V2:  $V_2 = (V_1 / T_1) T_2$

$$V_2 = (1.49 \text{ L} / 307.75 \text{ K}) (351.41 \text{ K}) = 1.70 \text{ L}$$

3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?

The combined gas law won't work here. You only have one condition and nothing changes.

$PV = nRT$  will work. ( $n$  = moles and  $R = 0.08026 \text{ L atm/K mol}$ )

$P = 1.04 \text{ atm}$ ,  $V = ? \text{ L}$ ,  $n =$  OOPS we need to convert from grams to moles!,  $T = 84.7 \text{ degrees Celsius} + 273 = 357.7 \text{ K}$

$n$ : Hydrogen gas is diatomic!

$$8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$$

$$\text{Isolate V: } V = (nRT) / P = (4.193069307 \text{ mol H}_2 \times 0.08026 \text{ L atm/K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$$

4. What volume is occupied by 56.75 g of oxygen gas at STP?

# 2nd Semester Final Chemistry High School

**PT Brinkman**

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The Nebraska Educational Journal, 1927   **Tests in Print** Oscar Krisen Buros, 1974   Phi Delta Kappan, 1928  
**American Machinist**, 1921   **Official Gazette** Philippines, 1999   Tests in Print II Oscar Krisen Buros, 1974  
**Annual Circular of the Illinois Industrial University** University of Illinois (Urbana-Champaign campus), 1929  
**Journal of the American Medical Association**, 1913 Includes proceedings of the association papers read at the annual sessions and lists of current medical literature   The Detroit Educational Bulletin Detroit (Mich.). Board of Education, 1924  
Vols 2-7 contain also Special bulletins pub during the same period   **Social Studies Tests and Reviews** Oscar Krisen Buros, 1975 Social Science Tests and Reviews consisting of the social science sections of the first seven MMYs and Tests in Print II includes 166 original test reviews written by 72 specialists five excerpted test reviews 71 references on the construction use and validity of specific tests a bibliography on in print social science tests references for specific tests cumulative name indexes for specific tests with references a publishers directory title index name index and a scanning index The 85 tests covered fall into the following categories 22 general 5 contemporary affairs 10 economics 7 geography 24 history 13 political science and 4 sociology   **Journal of Chemical Education**, 1925 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers  
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