

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from pH and pOH , and vice versa from pH and pOH .
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pH | 6. pOH |

Key Equation

- | | |
|-------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

Part A Completion

Read the statement and fill in the blank with the word or phrase that best completes the statement.

1. When a substance ionizes, it releases hydrogen ions. The concentration of H^+ in the autoionization of water is 1.0×10^{-7} moles per liter at 25°C and is equal to 1.0×10^{-7} moles per liter.
2. The pH of a solution is the negative logarithm of the concentration of H^+ in the solution. The pH of a solution is $-\log [\text{H}^+]$.
3. The pH of a solution is the negative logarithm of the concentration of H^+ in the solution. The pH of a solution is $-\log [\text{H}^+]$.
4. The pH of a solution is the negative logarithm of the concentration of H^+ in the solution. The pH of a solution is $-\log [\text{H}^+]$.
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**Kevin T. Patton, Frank B. Bell, Terry
Thompson, Peggie L. Williamson**

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