

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. pH | |

Key Equation

- | | |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include correct units when appropriate.

1. Water autoionizes. K_w is the ionization constant of water. K_w is the concentration of H_3O^+ and OH^- in pure water at $25^\circ C$.
2. Hydrochloric acid dissociates from HCl to H^+ and Cl^- . K_a is the ionization constant of HCl . K_a is the concentration of H^+ and Cl^- in pure HCl solution.
3. K_b is the ionization constant of a base. K_b is the concentration of OH^- and the conjugate acid of the base in pure K_b solution.
4. K_a is the ionization constant of an acid. K_a is the concentration of H^+ and the conjugate base of the acid in pure K_a solution.

5. K_w
6. 1.0×10^{-14}
7. 1.0×10^{-7}
8. 1.0×10^{-7}
9. 1.0×10^{-7}
10. 1.0×10^{-7}
11. 1.0×10^{-7}
12. 1.0×10^{-7}
13. 1.0×10^{-7}
14. 1.0×10^{-7}
15. 1.0×10^{-7}
16. 1.0×10^{-7}

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Wilhelm Mansfield Clark



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