

Approach to Arterial Blood Gases (ABGs)

1. Adequate oxygenation?

Normal ABG on Room Air = 7.40 / 40 / 90 / 24

Normal pCO_2 should be $\text{pCO}_2 \approx 4-5$ (pCO_2 in room air = 30)

$\text{pCO}_2 < 2$, then expected

Possible hypoxemia

2. Define the acid-base disturbance

$\text{pH} < 7.40 =$ Acidosis

$\text{pH} > 7.40 =$ Alkalosis

See slide on
Hypoxemia, Pathogenesis and Clinical Findings

3. Identify the primary (Δ^+) process

4. Compensation mechanisms in lungs and kidneys attempt to keep/retain CO_2 or HCO_3^- to maintain normal pH. Determination of compensation is appropriate (appropriate normal ratio of change in pCO_2 : HCO_3^- is shown in green boxes). If inappropriate, there is a secondary (Δ^-) process accounting

$\Delta^+ \text{ HCO}_3^-$

$\Delta^+ \text{ HCO}_3^-$

$\Delta^- \text{ HCO}_3^-$

5. Calculate anion gap (AG) to determine the presence and type of metabolic acidosis. This is done regardless of the Δ^+ or Δ^- process, as there may also be a hidden metabolic acidosis

Calculate $\text{AG} = \text{Na}^+ - \text{Cl}^- - \text{HCO}_3^-$ (normal is ≈ 12)

6. HAGMA: refer to normal blood buffer system, acid gain should match bicarbonate loss. If not, identify if another process is causing an inappropriate loss or gain of HCO_3^- .
 $\text{Gain of acid (AG)} = \text{AG} - 12$
 $\text{Loss of } \text{HCO}_3^- (\Delta \text{HCO}_3^-) = 24 - \text{HCO}_3^-$

Classically, the compensating anion is Cl^- , maintaining a normal AG

If a nonbicarbonate anion is present, $\Delta^- \text{ HCO}_3^-$ is compensated by Δ^+ other serum anions to maintain neutral charge

Where the cause is addition of an abnormal acid, HCO_3^- is used up as a buffer, and the compensating anion is the abnormal acid's conjugate base – which is not measured in the AG calculation

$\Delta^- \text{ HCO}_3^-$ is compensated by unmeasured anions ($\Delta^- \text{ the AG} - \text{f}$)

7. Calculate Cation Gap (net positive) or Urine net charge (for HAGMA) to narrow the etiology

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*With or some of publication

$\text{AG} = 12$

$\text{AG} > 12$

If no metabolic acidosis was identified in steps 2-4:
 No metabolic acidosis present

If a Δ^+ or Δ^- metabolic acidosis was identified in steps 2-4:
 Normal Anion Gap Metabolic Acidosis (NAGMA)

In this scenario, a normal total cation should be exceeding excess acid

High Anion Gap Metabolic Acidosis (HAGMA)

Calculate $\text{AG} = \text{Na}^+ - \text{Cl}^-$

Calculate $\text{pHCO}_3^- = 24 - \text{HCO}_3^-$

$\text{Na}^+ - \text{Cl}^- = \text{AG}$

$\text{Na}^+ - \text{Cl}^- = \text{AG}$

$\text{Na}^+ - \text{Cl}^- > \text{AG}$

$\text{Na}^+ - \text{Cl}^- > \text{AG}$

Loss of $\text{HCO}_3^- > \text{Gain of acid}$

Loss of $\text{HCO}_3^- > \text{Gain of acid}$

Additional process causing further loss of HCO_3^-

Only acid gain from HAGMA-causing HCO_3^- loss

$\text{HAGMA} + \text{NAGMA}$

$\text{HAGMA} + \text{Metabolic acidosis}$

$\text{HAGMA} + \text{Metabolic acidosis}$

$\text{HAGMA} + \text{Metabolic acidosis}$

Toxic alcohol ingestion (Δ^- serum bicarbonate and H^+ from metabolites produced), so cationality helps identify etiology

$\text{HAGMA} + \text{Metabolic acidosis}$

Calculate anion gap to determine cause of HAGMA

$\text{HAGMA} + \text{Metabolic acidosis}$

$= (\text{measured serum cation} - \text{measured serum anion}) + (\text{measured serum cation} - 12.5\text{mEq}) + \text{Urea} + \text{Serum glucose}$

$\text{HAGMA} + \text{Metabolic acidosis}$

$\text{Cation gap} > 10$ (metabolic present)

$\text{Cation gap} > 10$ (metabolic)

Toxic alcohol poisoning

Toxic alcohol poisoning

Test for other causes

Test for other causes

See slide on Normal Anion Gap Metabolic Acidosis: Pathogenesis and Laboratory Findings

See slide on High Anion Gap Metabolic Acidosis: Pathogenesis and Laboratory Findings

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**Marc N. Potenza, Kyle Faust, David
Faust**

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,1996

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