

# Meteorology Unit Study Guide

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Figure 1. The effect of the concentration of the solution on the adsorption of the dye.

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**Abstract**

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**Abstract** The authors examined the effects of a 12-week, 1000 kcal energy deficit diet on the body composition and metabolic profile of 10 obese women. The diet was composed of 25% protein, 40% carbohydrate, and 35% fat. The subjects lost 10.5% of their initial body weight and 15.5% of their initial fat mass. The diet also resulted in a 10% reduction in total energy expenditure, a 15% reduction in resting energy expenditure, and a 10% reduction in thermogenesis. The authors conclude that a 12-week, 1000 kcal energy deficit diet is an effective means of reducing body weight and fat mass in obese women.

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**Abstract** The purpose of this study was to determine the effect of a 12-week training program on the physical fitness and health-related quality of life (HRQL) of sedentary, middle-aged men. The study was a randomized, controlled trial. The intervention group (IG) consisted of 20 men who participated in a 12-week training program. The control group (CG) consisted of 20 men who did not participate in the training program. The IG showed significant improvements in physical fitness and HRQL compared to the CG. The results of this study suggest that a 12-week training program can improve physical fitness and HRQL in sedentary, middle-aged men.

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web 33mol o<sub>2</sub> 6mol br<sub>2</sub> 11mol o<sub>2</sub> 18mol br<sub>2</sub> identify the limiting reactant s and excess reactant s the limiting reactant is o<sub>2</sub> since it would yield the least amount of product 18 mol br<sub>2</sub> the excess reactant is c<sub>2</sub>h<sub>3</sub>br<sub>3</sub> since its complete reaction would have yielded up to 54 mol br<sub>2</sub> think about your result

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web step 4 the reactant that produces a smaller amount of product is the limiting reagent mg mg produces less mgo mgo than does o<sub>2</sub> o<sub>2</sub> 3 98 g mgo vs 25 2 g mgo therefore mg is the limiting reagent in this reaction step 5 the reactant that

produces a larger amount of product is the excess reagent

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web if you start with 14 82 g of ca oh 2 c a o h 2 and 16 35 g of h<sub>2</sub>so<sub>4</sub> h 2 s o 4 a determine the limiting reagent b determine the number of moles of h<sub>2</sub>o h 2 o produced c determine the number of grams of caso<sub>4</sub> c a s o 4 produced d determine the number of grams of excess reagent left 1 make sure the equation is balanced

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