

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. acid rain | 6. acid rain |
| 7. autoionization | 8. acid rain | 9. pH |

Key Equation

- | | |
|------------------------------------|------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}^+]$ |

Part A Completion

Write the equation for the autoionization of water. What is the concentration of H^+ in pure water? What is the concentration of OH^- in pure water?

What is the pH of a solution with $[\text{H}^+] = 1.0 \times 10^{-7} \text{ M}$?

What is the pH of a solution with $[\text{OH}^-] = 1.0 \times 10^{-7} \text{ M}$?

What is the pH of a solution with $[\text{H}^+] = 1.0 \times 10^{-7} \text{ M}$?

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What is the pH of a solution with $[\text{OH}^-] = 1.0 \times 10^{-7} \text{ M}$?

What is the pH of a solution with $[\text{H}^+] = 1.0 \times 10^{-7} \text{ M}$?

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

2. $1.0 \times 10^{-7} \text{ M}$

3. $1.0 \times 10^{-7} \text{ M}$

4. $1.0 \times 10^{-7} \text{ M}$

5. $1.0 \times 10^{-7} \text{ M}$

6. $1.0 \times 10^{-7} \text{ M}$

7. $1.0 \times 10^{-7} \text{ M}$

8. $1.0 \times 10^{-7} \text{ M}$

9. $1.0 \times 10^{-7} \text{ M}$

10. $1.0 \times 10^{-7} \text{ M}$

11. $1.0 \times 10^{-7} \text{ M}$

12. $1.0 \times 10^{-7} \text{ M}$

19 2 Hydrogen Ions And Acidity Answer Key

Frank R. Foulkes



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