

Chemistry I

Study Guideline – Unit Three

Atomic Structure

By the end of this chapter, you should be able to:

1. Describe the nature and significance of scientific modeling in general and several of the atoms.
2. Compare and Contrast the models of the atom including Dalton's, Thomson's and Rutherford's.
3. Explain how atoms are related to electricity.
4. Describe how J. J. Thomson investigated the charge-to-mass ratio of an electron and a proton using a CRT.
5. Relate the nature of radioactivity to Ernst Rutherford's alpha scattering experiment and how it showed the existence of the nucleus.
6. Identify the characteristics of Alpha, Beta and Gamma radiation.
7. Name and describe three subatomic particles.
8. Define atomic number, mass number, isotopes and ions.
9. Define and Explain (average) atomic mass.
10. Determine the number of protons, neutrons and electrons in an isotope from its name or symbol.
11. Determine the atomic mass of an element, given the masses and percent abundance of the elements isotopes.
12. Explain why elements above atomic number 83 are radioactive.
13. Describe how a mass spectrometer is used to determine the relative masses and abundances of isotopes.

Atomic Structure Study Guide Answers

Marc Loudon, Jim Parise

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