

Chemical Reactions:

1. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. What is the ΔH°_f for the reaction?

A) greater than 2.00 kcal/mol
 B) greater than 1.00 kcal/mol
 C) less than 1.00 kcal/mol
 D) less than 2.00 kcal/mol
 E) greater than 4.00 kcal/mol

2. A 1.0-mol sample of PCl_3 reacts with excess oxygen according to the equation: $2\text{PCl}_3(\text{g}) \rightarrow 2\text{PCl}_5(\text{g}) + \text{O}_2(\text{g})$. Below, which reaction will yield the greatest amount of O_2 after PCl_3 is removed?

A) 0.0 mol
 B) 1.0 mol
 C) 1.2 mol
 D) 1.7 mol
 E) 2.0 mol

3. The reaction of NO_2 is step characterized by the equation: $4\text{NO}_2 \rightarrow 3\text{O}_2 + 2\text{N}_2\text{O}_5$. If 1.0 mol of NO_2 is added to 1.0 mol of N_2O_5 and the reaction is allowed to proceed, how will the reaction rate change?

A) 1.000 mol/d
 B) 0.750 mol/d
 C) 1.000 mol/d
 D) 1.50 mol/d
 E) 2.00 mol/d

4. Calculated are the following reactions, or from the experiments, are unknowns, calculated.

$\text{H}_2\text{O}_2(\text{l}) + \text{H}_2\text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O}_2(\text{l}) = 300\text{J/g}$

$\text{H}_2 + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$ (enthalpy of difference, ΔH°_f for production of 1.00 mol of H_2O_2 , which is the difference of 200)

A) O_2
 B) H_2
 C) H_2
 D) O_2
 E) C

5. Under the following conditions, $\text{NO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightarrow 2\text{NO}(\text{g}) + \text{O}_2(\text{g})$ is

higher energy molecules of NO_2 are produced in?

A) 1.000 mol/d
 B) 1.000×10^3
 C) 1.000×10^6
 D) 1.000
 E) same as above

6. The following reaction $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$ is performed. Below, which number of molecules would be produced from 1.000 mol of hydrogen and excess oxygen?

A) 1.000 mol/d
 B) 2.000 mol/d
 C) 0.500 mol/d
 D) 0.500 mol
 E) 0.250 mol

7. Under the following conditions, $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}(\text{g}) + \text{H}_2\text{O}(\text{l})$. What mass of oxygen, $\text{O}_2(\text{g})$ is required to react completely with 1.00 g of $\text{CH}_4(\text{g})$?

A) 1.000×10^3 g
 B) 1.000×10^4 g
 C) 4.000×10^3 g
 D) 4.000×10^4 g
 E) 0.1000×10^3 g

8. Under reactor conditions and quantities of reactants, the following reaction of CH_4 with propane is carried out. How many molecules of CH_4 are required to react completely with 1.000 mol of propane?



A) 1.000×10^{10} molecules
 B) 1.000×10^{11} molecules
 C) 4.000×10^{10} molecules
 D) 8.000×10^{10} molecules
 E) 7.000×10^{10} molecules

9. Calculate the mass of carbon dioxide produced from 1.00 g of octane, C_8H_{18} , in the following reaction.



A) 7.00×10^{-3} g
 B) 1.000×10^{-3} g
 C) 9.00×10^{-3} g
 D) 1.200×10^{-3} g
 E) 1.600×10^{-3} g

10. Calculate the molecules of oxygen required to react with 1.00 g of sulfur in the following reaction.



A) 1.000×10^{10} molecules O_2
 B) 1.000×10^{11} molecules O_2
 C) 1.71×10^{10} molecules O_2
 D) 2.40×10^{10} molecules O_2
 E) 3.12×10^{10} molecules O_2

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Cambridge Checkpoints VCE Chemistry Units 3 and 4 2013 Roger Slade, Maureen Slade, 2012-12-14 Cambridge Checkpoints VCE are updated regularly to provide you with the most up to date exam preparation available **Cracking the AP Chemistry Exam, 2012 Edition** Paul Foglino, 2011-08-02 Provides techniques for achieving high scores on the AP chemistry exam and includes two full length practice tests *Roadmap to the Regents* Sasha Alcott, 2003 If Students Need to Know It It's in This Book This book develops the chemistry skills of high school students It builds skills that will help them succeed in school and on the New York Regents Exams Why The Princeton Review We have more than twenty years of experience helping students master the skills needed to excel on standardized tests Each year we help more than 2 million students score higher and earn better grades We Know the New York Regents Exams Our experts at The Princeton Review have analyzed the New York Regents Exams and this book provides the most up to date thoroughly researched practice possible We break down the test into individual skills to familiarize students with the test's structure while increasing their overall skill level We Get Results We know what it takes to succeed in the classroom and on tests This book includes strategies that are proven to improve student performance We provide a breakdown of the skills based on New York standards and objectives hundreds of practice questions organized by skill two complete practice New York Regents Exams in Physical Setting Chemistry *Master the CLEP* Peterson's, 2012-04-30 Peterson's Master the CLEP is the essential prep guide for non traditional students who are looking to begin a college career and for current students who want to save on tuition costs by testing out of certain courses When you use this guide you will get a basic overview of the CLEP a look at the general examinations and the vast number of subject specific exams tips for registering and preparing for the CLEP and a guide to interpreting your test scores This is followed by extensive reviews of the five CLEP general examinations with pre and post tests and subject matter overviews that will help you pinpoint your strengths and weaknesses The Appendix will go into detail about a number of CLEP subject exams *CLEP Success* Peterson's, 2007-05-04 Offers advice on improving scores on the CLEP and includes reviews for five CLEP subjects and five full length practice tests with explanatory answers

Journal of Chemical Education, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers *Tests* Richard C. Sweetland, Daniel J. Keyser, 1983 Descriptions of over 3000 tests in English intended as a guide for psychologists educators and other personnel who need test information to meet their assessment needs Topical arrangement under 3 main sections of psychology education and business Each entry gives test name associated personal names intended age group purpose description time range scoring cost and publisher Indexes by test titles authors publishers visually impaired tests and scoring services *The ETS Test Collection Catalog* Educational Testing Service. Test Collection, 1993 The major source of information on the availability of standardized tests Wilson Library Bulletin Covers commercially available standardized tests and hard to locate

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