



## Chapter 14 – DNA: The Genetic Material

Study questions from Chapter 14 of the textbook, *Principles of Biology*

1. What were the major pieces for DNA that were identified?

Chromosomes, DNA, and RNA

2. What are the components of a nucleic acid?

nitrogen, phosphate groups, nitrogenous bases



3. What is the difference between deoxyribonucleic acid and ribonucleic acid?

deoxyribonucleic acid (DNA) contains deoxyribose nucleic acid (RNA) contains ribose

4. What is a homopolymer? Is RNA a homopolymer?

homopolymer: same nucleotide repeating



5. What is a homopolymer? Is RNA a homopolymer?

homopolymer: same nucleotide repeating



6. What is a homopolymer? Is RNA a homopolymer?

homopolymer:



7. What is a homopolymer? Is RNA a homopolymer?

homopolymer:



8. What is a homopolymer? Is RNA a homopolymer?

homopolymer: same nucleotide repeating

# Biology The Genetic Material Study Guide Answers

**William K. Purves, Edward  
Dzialowski, Lindsay Goodloe, Betty  
McGuire, Nancy Guild, Paula Mabee**

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web correct answer c both a and b q2 the density of 1 m solution of hcl is 1 0585 g ml the molality of the solution is 1 0585 1 0 10 0 0585 correct answer b 1 explanation molality begin array 1 m left frac 1000 m 1000d mm right end array

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web sep 13 2023 the only difference between molality and molarity is in the denominator molality is measured in kilograms of solvent whereas molarity is measured in liters of solution when studying the properties of solutions related to vapour

pressure and temperature changes concentrations expressed in molality are used

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web molarity is the number of moles of a substance per litre of solution also known as molar concentration a capital m signifies solutions labelled with molar concentration a 1 0 m solution contains 1 mole of solute per litre of solution molality is the number of solute moles per kilogram

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web 1 molarity mv mass molar mass x 0 4500 1 0 825 g 141 9579 g mol x 0 0129 m 2 molality 0 825 g 141 9579 g mol 0 00581158 mol 0 00581158 mol 0 4500 kg 0 0129 m 3 mole fraction na 2 hpo 4 0 825 g 141 9579 g mol 0 00581158 mol h 2 o 450 0 g 18 015 g mol 24 97918401 mol

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web apr 7 2014 molarity moles of solute liters of solution 8 4 2 2 a first convert 250 ml to liters 250 1000 0 25 then calculate molarity 5 moles 0 25 liters 20 m 3 c a solution with molarity 2 requires 2 m of n a oh per liter so 4 x 2 8 m 4 a a solution of molarity 1 5 m requires 1 5 mol of na to every litre of solvent

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web determine the molarity for each of the following solutions 0 444 mol of  $\text{CoCl}_2$  in 0 654 l of solution 98 0 g of phosphoric acid  $\text{H}_3\text{PO}_4$  in 1 00 l of solution 0 2074 g of calcium hydroxide  $\text{Ca(OH)}_2$  in 40 00 ml of solution 10 5 kg

*16 11 molality chemistry libretexts - Aug 15 2023*

web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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