



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">hydrogen ionacidbasestrong acidweak acidpH	<ul style="list-style-type: none">hydrogen ion concentrationacidicbasicneutraltitrationtitration curve
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

Key Diagrams

- Titration curve
- Titration curve

Points of Consideration

Acid dissociation constant is a measure of the strength of an acid. The greater the acid dissociation constant, the stronger the acid.

Strong acids dissociate completely in aqueous solution. The concentration of these acids is equal to the concentration of their ions.

Strong acids dissociate completely in aqueous solution. The concentration of these acids is equal to the concentration of their ions.

Weak acids dissociate partially in aqueous solution. The concentration of these acids is greater than the concentration of their ions.

Weak acids dissociate partially in aqueous solution. The concentration of these acids is greater than the concentration of their ions.

Key Equations

$\text{pH} = -\log_{10}[\text{H}^+]$	$\text{pH} = \log_{10}[\text{OH}^-]$
$\text{pH} = \log_{10}[\text{H}^+]$	$\text{pH} = \log_{10}[\text{OH}^-]$
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