

## Acid and Base Strength Worksheet

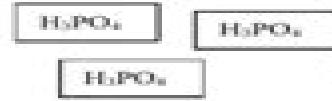
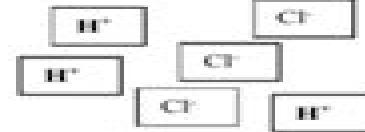
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What makes an acid **acidic** are the excess **Hydrogen ions ( $H^+$ )**\* that these substances cause in a water solution.

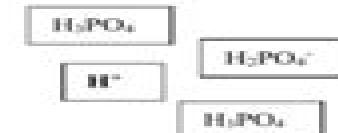
**Strong acids, like  $HCl$ , dissociate (break up) completely** in solution, giving a maximum number of hydrogen ions. Another way to say this is, strong acids **IONIZE** completely while weak acids don't.



$HCl$  particles ionize completely in water to yield  $H^+$  and  $Cl^-$ .  
Thus  $HCl$  is a **STRONG ACID**.



Only some  $H_3PO_4$  particles ionize in water. So in a solution there will be intact  $H_3PO_4$  particles as well as  $H^+$  and  $H_2PO_4^-$ . Thus  $H_3PO_4$  is a **WEAK ACID**.



So, how do you know which acids are strong and which are weak?  
Easy: there are a few rules.

First, a few definitions:

- A **binary acid** is made of **two elements**: hydrogen and something else.  $HCl$  is a binary acid made of hydrogen and chloride. Count 'em. Hydrogen (one), chloride (two). Ta-da!
- A **ternary acid** is made of **more than two elements**: hydrogen and at least two other elements.  $H_2SO_4$ , with three elements, is a ternary acid. So is  $H_3PO_4$  and acetic acid,  $HC_2H_3O_2$ . (Notice acid formulas always start with a hydrogen. These are the hydrogens that can break free in water to a lesser or greater degree depending on the strength of the acid.)

### THE RULES for deciding if an acid is strong or weak!

- **The only strong binary acids are  $HCl$ ,  $HBr$ , and  $HI$ .** (Notice Cl, Br, and I are together on the periodic table.) All other binary acids are weak.
- Ternary acids are a little trickier, but not much. Compare the number of oxygens and hydrogens in the formula. **If the number of oxygens exceeds the number of hydrogens by 2 or more, the acid is strong.** All other ternary acids are weak.

$H_3PO_4$       3 hydrogens and 4 oxygens (only one more oxygen than hydrogen). **WEAK**

$H_2SO_4$       2 hydrogens and 4 oxygens (two more oxygens than hydrogen). **STRONG**.

\*Technically, the free  $H^+$  usually combines with a water molecule to form a hydronium ion ( $H_3O^+$ ), but we will ignore this for the time being.

## 19 2 Hydrogen Ions And Acidity Worksheet Answers

**W. M. Clark**

## 19 2 Hydrogen Ions And Acidity Worksheet Answers:

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