

Name : \_\_\_\_\_ Date : \_\_\_\_\_

## Acids and Bases Problems

- Without doing any calculations, identify each of the following solutions as acidic or basic:
  - $[\text{H}_3\text{O}^+] = 3.4 \times 10^{-8} \text{ M}$  \_\_\_\_\_
  - $[\text{OH}^-] = 1.8 \times 10^{-5} \text{ M}$  \_\_\_\_\_
  - $[\text{OH}^-] = 6.5 \times 10^{-12} \text{ M}$  \_\_\_\_\_
  - $[\text{H}_3\text{O}^+] = 2.6 \times 10^{-4} \text{ M}$  \_\_\_\_\_
- A solution of HBr has  $[\text{H}_3\text{O}^+] = 4.5 \times 10^{-3} \text{ M}$ . Calculate the pH of this solution.

- Complete the following table.

pH	$[\text{H}_3\text{O}^+]$	$[\text{OH}^-]$	Acidic, basic, or neutral?
	$5.4 \times 10^{-4}$		
	$1.3 \times 10^{-5}$		
10.75	$1.8 \times 10^{-11}$	$5.6 \times 10^{-4}$	

- If the pH of a solution is 10.3, what is the  $[\text{H}^+]$  concentration?
- If the  $[\text{H}^+]$  is  $2.1 \times 10^{-12} \text{ M}$   $\text{HClO}_4$ , what is the pH? Is the solution ACIDIC, BASIC, or NEUTRAL?
- Calculate the pH if the  $[\text{OH}^-]$  concentration is  $5.9 \times 10^{-1} \text{ M}$ ? Is the solution ACIDIC, BASIC, or NEUTRAL?

# Acids And Bases Practice Worksheer

**Yuli Rahmawati, Peter Charles Taylor**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

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