

10. As you saw in Model 1, all Arrhenius bases in Model 1 have an  $\text{OH}^-$  ion in their chemical formulas. Write a balanced chemical reaction for the reaction of  $\text{HCl(aq)}$  with  $\text{OH}^-(\text{aq})$  to illustrate that the hydroxide ion is also a Brønsted-Lowry base.



11. If you reverse Reaction 1 in Model 2, the following reaction is obtained.



- a. For the reaction above, which reactant is acting like a Brønsted-Lowry acid? How can you tell?

$\text{H}_3\text{O}^+$  because it donates a proton

- b. For the reaction above, which reactant is acting like a Brønsted-Lowry base? How can you tell?

$\text{Cl}^-$  because it accepts a proton

12. Write the reverse reactions for Reactions 2 and 3 in Model 2. Label the Brønsted-Lowry acid and base reactants for each reaction.



Reaction 3



### Model 3 – Conjugate Acid–Base Pairs



13. All acid–base reactions have two conjugate acid–base pairs. One conjugate acid–base pair in the reaction in Model 3 is  $\text{H}_3\text{O}^+/\text{H}_2\text{O}$ . List the other acid–base pair in the reaction.



14. Why is  $\text{HCO}_3^-$  considered the “acid” part of the pair in the reaction in Model 3?

It donates a proton

15. Why is  $\text{CO}_3^{2-}$  considered the “base” part of the pair in the reaction in Model 3?

It accepts a proton

16. The “Read This” box before Model 2 calls the transfer of a  $\text{H}^+$  ion a “proton transfer.” Explain why “ $\text{H}^+$  ion” and “proton” are synonymous.

Hydrogen is just a proton and an electron, so when it is + (1 electron removed) it is just a proton

# Acids And Bases Pogil Answer Key

**L Darling-Hammond**



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