



10. As you saw in Model 1, all Arrhenius bases in Model 1 have an OH⁻ ion in their chemical formulas. Write a balanced chemical reaction for the reaction of HCl(aq) with OH⁻(aq) to illustrate that the hydroxide ion is also a Brønsted-Lowry base.



11. If you reverse Reaction 1 in Model 2, the following reaction is obtained.



a. For the reaction above, which reactant is acting like a Brønsted-Lowry acid? How can you tell?

H_2O^+ because it donates a proton

b. For the reaction above, which reactant is acting like a Brønsted-Lowry base? How can you tell?

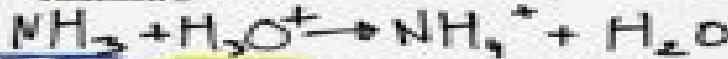
Cl^- because it accepts a proton

12. Write the reverse reactions for Reactions 2 and 3 in Model 2. Label the Brønsted-Lowry acid and base reactants for each reaction.

Reaction 2



Reaction 3



Model 3 – Conjugate Acid–Base Pairs



13. All acid–base reactions have two conjugate acid–base pairs. One conjugate acid–base pair in the reaction in Model 3 is $\text{H}_2\text{O}/\text{H}_3\text{O}^+$. List the other acid–base pair in the reaction.



14. Why is HCO_3^- considered the "acid" part of the pair in the reaction in Model 3?

It donates a proton

15. Why is CO_3^{2-} considered the "base" part of the pair in the reaction in Model 3?

It accepts a proton

16. The "Read This" box before Model 2 calls the transfer of a H⁺ ion a "proton transfer." Explain why "H⁺ ion" and "proton" are synonymous.

Hydrogen is just a proton and an electron, so when it is + (1 electron removed) it is just a proton

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