

1. Bases react with acids to produce water and salt and this type of reaction is called a neutralization reaction.

2. What is the chemical name of H_2S ? Hydrosulfuric acid $\text{ide} \rightarrow \text{hydro-ic}$

3. What is the chemical name of H_2SO_3 ? Sulfurous acid $\text{ite} \rightarrow \text{ous}$

4. What is the chemical name of H_2SO_4 ? Sulfuric acid $\text{ate} \rightarrow \text{ic}$

5. What is the chemical formula for hydraphosphoric acid? H_3P

6. What is the chemical formula for phosphoric acid? H_3PO_4

7. What is the chemical name of $\text{Ca}(\text{OH})_2$? calcium hydroxide

8. Arrhenius Acids & bases ionize to form H^+ and OH^- respectively is an assumption of what theory?

9. Bronsted-Lowry acids donate a proton while bases accept a proton is an assumption of what theory?
donate accept

10. Circle the Bronsted-Lowry acid and box the Bronsted-Lowry base in the following reaction.



11. A student places a 0.34 gram sample of sodium in an acidic solution and a violent reaction occurs producing what gas? hydrogen gas

Label the following as an Arrhenius Acid or an Arrhenius Base

12. acid HCl
13. base NaOH
14. base $\text{Mg}(\text{OH})_2$
15. acid H_2SO_4

Label the following reactions as:

- precipitation (& circle the precipitate)
- neutralization or

• neutralization of

16. neutralization acid HNO_3 + base $\text{NaOH} \rightarrow$ salt NaNO_3 + water H_2O

17. neutralization $\text{HCl} + \text{KOH} \rightarrow \text{H}_2\text{O} + \text{KCl}$

18. precipitate $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$

Acid Base Study Guide Answer Key

Michael B. Smith

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