

1. Bases react with acids to produce water and salt and this type of reaction is called a neutralization reaction.
2. What is the chemical name of H_2S ? hydrosulfuric acid ide \rightarrow hydro-ic
3. What is the chemical name of H_2SO_3 ? sulfurous acid ite \rightarrow ous
4. What is the chemical name of H_2SO_4 ? sulfuric acid ate \rightarrow ic
5. What is the chemical formula for hydrophosphoric acid? H_3P
6. What is the chemical formula for phosphoric acid? H_3PO_4
7. What is the chemical name of $\text{Ca}(\text{OH})_2$? calcium hydroxide
8. Arrhenius Acids & bases ionize to form H^+ and OH^- respectively is an assumption of what theory?
9. Bronsted-Lowry acids donate a proton while bases accept a proton is an assumption of what theory?
10. Circle the donate Bronsted-Lowry acid and box the accept Bronsted-Lowry base in the following reaction.



11. A student places a 0.34 gram sample of sodium in an acidic solution and a violent reaction occurs producing what gas? hydrogen gas

Label the following as an Arrhenius Acid or an Arrhenius Base

12. acid HCl H^+ -
13. base NaOH - OH^-
14. base $\text{Mg}(\text{OH})_2$
15. acid H_2SO_4

Label the following reactions as:

- precipitation (& circle the precipitate)
- neutralization or



Acid Base Study Guide Answer Key

Michael B. Smith



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