

Chemical Reactions:

1. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. What is the ΔH°_f for O_2 in kJ/mol ?

A) greater than 2.0 kJ/mol O_2
 B) greater than 1.0 kJ/mol O_2
 C) equal to 1.0 kJ/mol O_2
 D) less than 1.0 kJ/mol O_2
 E) greater than 4.0 kJ/mol O_2

2. A 1.0-molar sample of H_2SO_4 reacts with excess CaCO_3 according to the equation:

$$2\text{H}_2\text{SO}_4(\text{aq}) \rightarrow 2\text{H}_2\text{O} + \text{SO}_2(\text{g})$$
 Below, which samples of O_2 are consumed in the reaction?

A) 0.0 mol
 B) 1.0 mol
 C) 1.2 mol
 D) 1.7 mol
 E) 2.0 mol

3. The reaction of H_2 with O_2 is represented by the equation: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$. If 1.0 mol of H_2 reacts with 1.0 mol of O_2 , which of the following is true about the reaction?

A) 1.0 mol of H_2O is produced.
 B) 1.0 mol of H_2O is consumed.
 C) 1.0 mol of H_2O is formed.
 D) 1.0 mol of H_2O is consumed.
 E) 0.5 mol of H_2O is formed.

4. Calculated are the following reactions, or both?

Requirements are molarities, not amounts.
 $\text{H}_2\text{O}_2(\text{aq}) + \text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}_2(\text{aq}) = 2\text{H}_2\text{O}_2(\text{g})$

If 1.0 g of O_2 is added to excess H_2O_2 , which is the following reaction?

A) O_2
 B) H_2O_2
 C) H_2
 D) H_2O
 E) C

5. Rather than the following reaction, w hours CO_2 requirements are molarities, not amounts.
 $\text{CH}_3\text{CH}_2\text{OH} + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O} + \text{CO}_2(\text{g})$

What mass of oxygen, $\text{O}_2(\text{g})$ is required to react completely with 1.0 L mol of $\text{CH}_3\text{CH}_2\text{OH}$?

A) $1.000 \times 10^3 \text{ g}$
 B) $2.000 \times 10^3 \text{ g}$
 C) 4.000 g
 D) 8.000 g
 E) 0.1000 g

6. The following reaction $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow$ $2\text{H}_2\text{O}$ shows molar masses of reactants. Would the product exceed 36.0 g/mol of hydrogen and oxygen?

A) 1.0 g/mol
 B) 2.0 g/mol
 C) 10.0 g/mol
 D) 20.0 g/mol
 E) 30.0 g/mol

7. Rather than the following reaction, w hours CO_2 requirements are molarities, not amounts.
 $\text{C}_2\text{H}_5\text{OH} + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O} + \text{CO}_2(\text{g})$

What mass of oxygen, $\text{O}_2(\text{g})$ is required to react completely with 1.0 L mol of $\text{C}_2\text{H}_5\text{OH}$?

A) $1.000 \times 10^3 \text{ g}$
 B) $2.000 \times 10^3 \text{ g}$
 C) 4.000 g
 D) 8.000 g
 E) 0.1000 g

8. Rather than the following reaction, w hours CO_2 requirements are molarities, not amounts.
 $\text{CH}_3\text{CH}_2\text{OH} + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O} + \text{CO}_2(\text{g})$

What mass of oxygen, $\text{O}_2(\text{g})$ is required to react completely with 1.0 L mol of $\text{CH}_3\text{CH}_2\text{OH}$?

A) $1.000 \times 10^3 \text{ g/mol}$
 B) $1.000 \times 10^2 \text{ g/mol}$
 C) $4.000 \times 10^2 \text{ g/mol}$
 D) $8.000 \times 10^2 \text{ g/mol}$
 E) 0.1000 g/mol

9. Calculated are the masses of carbon dioxide produced from 1.0 g of octane, C_8H_{18} , in the following reaction.

$$2\text{C}_8\text{H}_{18}(\text{g}) + 25\text{O}_2(\text{g}) \rightarrow 16\text{CO}_2(\text{g}) + 18\text{H}_2\text{O}(\text{g})$$

A) $7.00 \times 10^{-3} \text{ g CO}_2$
 B) $1.75 \times 10^{-3} \text{ g CO}_2$
 C) $9.00 \times 10^{-3} \text{ g CO}_2$
 D) $1.200 \times 10^{-3} \text{ g CO}_2$
 E) $1.800 \times 10^{-3} \text{ g CO}_2$

10. Calculated are the molarities of oxygen required to react with 1.0 L mol of sulfur in the following reaction.

$$2\text{S}(\text{s}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{SO}_2(\text{g})$$

A) $1.000 \times 10^3 \text{ g/mol}$ O_2
 B) $0.250 \times 10^3 \text{ g/mol}$ O_2
 C) $1.75 \times 10^2 \text{ g/mol}$ O_2
 D) $0.400 \times 10^3 \text{ g/mol}$ O_2
 E) $0.100 \times 10^3 \text{ g/mol}$ O_2

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Journal of Chemical Education, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers *Tests* Richard C. Sweetland, Daniel J. Keyser, 1983 Descriptions of over 3000 tests in English intended as a guide for psychologists educators and other personnel who need test information to meet their assessment needs Topical arrangement under 3 main sections of psychology education and business Each entry gives test name associated personal names intended age group purpose description time range scoring cost and publisher Indexes by test titles authors publishers visually impaired tests and scoring services

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