

## Semester 2 Final Exam Review – Answer Key

1. Mendeleev
2. Electrons in the outer-most energy shell of an atom
3. F, O, N
4. -ide
5. Atoms with a full valence shell
6. Molarity (M) = moles per liter = moles/liters
7. Basic, alkaline
8. Group or Family
9.  $6.02 \times 10^{23} \times 5 = 3.01 \times 10^{24}$
10. Fluid, compressible, take the shape and volume of the container, molecules are far apart (low density)
11. Period
12. Acidic
13. High boiling point, low vapor pressure
14. A homogeneous mixture of only one phase; solute dissolved in a solvent
15. 22.4L
16. Endothermic
17.  $\text{Zn}(\text{NO}_3)_2$
18. H, N, O, F, Cl, Br, I
19. Nature of the reactants, temperature, surface area, concentration, catalyst
20. Calorie = energy required to raise 1g of water 1 degree Celcius
21. Volume is directly proportional to temperature @ a constant pressure ( $V_1/T_1 = V_2/T_2$ )
22. Iron(II) chloride
23.  $\text{Ba}(\text{NO}_3)_2$
24.  $\text{PbCO}_3$
25. Reactants react to form products
26. Synthesis, decomposition, combustion, single replacement, double replacement
27. Water and carbon dioxide
28. Right side,  $-\Delta H$
29. The energy require to form an activated complex and have a reaction proceed to form products
30. Exothermic,  $\Delta H = -585\text{kJ}$ 
  - Convert mass to moles:  $34\text{g} + 17\text{g/mol} = 2 \text{ mol NH}_3$
  - Multiply by energy/mole ratio:  $2 \text{ mol NH}_3 \times (-1170\text{kJ})/4 \text{ mol NH}_3 = -585\text{kJ}$
31. Concentration, temperature, pressure (gas, heterogeneous system)
32. A buffer solution is a solution of a conjugate acid-base pair. A buffer system prevents significant changes in pH when a small amount of acid or base is added to the solution. Your body has a bicarbonate buffer system in the blood which helps maintain you blood's pH.
33. Unsaturated = more room for solute @ temp; saturated = solute maxed out @ temp; supersaturated = more solute than maximum for given temp

# 2nd Semester Final Exam Review Science Answer

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## **Table of Contents 2nd Semester Final Exam Review Science Answer**

1. Understanding the eBook 2nd Semester Final Exam Review Science Answer
  - The Rise of Digital Reading 2nd Semester Final Exam Review Science Answer
  - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Final Exam Review Science Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 2nd Semester Final Exam Review Science Answer
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2nd Semester Final Exam Review Science Answer
  - Personalized Recommendations
  - 2nd Semester Final Exam Review Science Answer User Reviews and Ratings
  - 2nd Semester Final Exam Review Science Answer and Bestseller Lists
5. Accessing 2nd Semester Final Exam Review Science Answer Free and Paid eBooks

- 2nd Semester Final Exam Review Science Answer Public Domain eBooks
- 2nd Semester Final Exam Review Science Answer eBook Subscription Services
- 2nd Semester Final Exam Review Science Answer Budget-Friendly Options
- 6. Navigating 2nd Semester Final Exam Review Science Answer eBook Formats
  - ePub, PDF, MOBI, and More
  - 2nd Semester Final Exam Review Science Answer Compatibility with Devices
  - 2nd Semester Final Exam Review Science Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 2nd Semester Final Exam Review Science Answer
  - Highlighting and Note-Taking 2nd Semester Final Exam Review Science Answer
  - Interactive Elements 2nd Semester Final Exam Review Science Answer
- 8. Staying Engaged with 2nd Semester Final Exam Review Science Answer
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 2nd Semester Final Exam Review Science Answer
- 9. Balancing eBooks and Physical Books 2nd Semester Final Exam Review Science Answer
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 2nd Semester Final Exam Review Science Answer
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 2nd Semester Final Exam Review Science Answer
  - Setting Reading Goals 2nd Semester Final Exam Review Science Answer
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 2nd Semester Final Exam Review Science Answer
  - Fact-Checking eBook Content of 2nd Semester Final Exam Review Science Answer
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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