

THE OFFICIAL STUDY GUIDE FOR: "CHEMICAL KINETICS"

Multiple Choice Section: This study guide is a compilation of questions from provincial exams since April 1994. I urge you to become intimately familiar with question types. You will notice that questions from one year to another are very similar in their composition. Identification of question types will allow you to be more efficient in answering these questions on the provincial examination. My recommendations for using this study guide are as follows:

- DO ALL THE QUESTIONS** in this booklet. These are actual Provincial Exam questions! Your own provincial exam and unit test will include questions similar to the ones in this booklet!
- RESIST THE URGE TO LOOK AT THE ANSWER KEY** until you have given all the questions in the section your best effort. Don't do one question, then look at the key, then do another and look at the key, and so on. Each time you look at one answer in the study guide, your eye will notice other answers around them, and this will reduce the effectiveness of those questions in helping you to learn.
- LEARN FROM YOUR MISTAKES!** If you get a question wrong, **figure out why!** If you are having difficulty, **talk to your study partner**, or maybe **phone someone in your Peer Tutoring group**. Get together with group members or other students from class and work on these questions together. Explain how you got your answers to tough questions to others. In explaining yourself to someone else, you will learn the material better yourself (try it!) Ask your teacher to explain the questions to you during tutorial or after school. **Your goal should be to get 100% on any Chemistry 12 multiple choice test** - learning from your mistakes in this booklet will really help you in your efforts to meet this goal!
- This is REALLY CRUCIAL: DO NOT mark the answer anywhere on the questions themselves.** For example, do not circle any of options A, B, C or D - instead use a different sheet of paper to place your answers on. By avoiding this urge, you can re-use this study guide effectively again, when preparing for your final exam. In the box to the left, put an asterisk or small note to yourself to indicate that you got the question wrong and need to come back to it. If you got the question correct initially, a check mark might be assurance that you understand this type of question and therefore can concentrate on other questions that present a challenge to you.
- Check off the STATUS box on the PRE-SCHEDULED LEARNING OUTCOMES sheet.** I have tried to organize the questions in the identical sequence to which they appear on your Acid-Base Prescribed Learning Outcome sheet. By doing this, you can be confident that you know everything you need to know for both the UNIT EXAM and PROVINCIAL EXAM!

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INTRODUCTION

- A2 The rate of a chemical reaction can be expressed in
 A. grams per mole. B. energy consumed per mole.
 C. volume of gas per unit time. D. moles formed per litre of solution.
- A2 Consider the following reaction:
 Under certain conditions, the rate of decomposition of NO_2 is $3.2 \times 10^{-3} \text{ mol/s}$. The rate of formation of O_2 is
 A. $1.6 \times 10^{-3} \text{ mol/s}$
 B. $3.2 \times 10^{-3} \text{ mol/s}$
 C. $4.8 \times 10^{-3} \text{ mol/s}$
 D. $6.4 \times 10^{-3} \text{ mol/s}$
- A2 Consider the following reaction:

$$2\text{N}_2\text{O}_5(\text{g}) \rightarrow 4\text{NO}_2(\text{g}) + \text{O}_2(\text{g})$$

 At a certain temperature the rate of decomposition of N_2O_5 is $2.5 \times 10^{-6} \text{ mol/s}$. The rate of formation of NO_2 is:
 A. $1.0 \times 10^{-6} \text{ mol/s}$ B. $1.3 \times 10^{-6} \text{ mol/s}$ C. $2.5 \times 10^{-6} \text{ mol/s}$ D. $5.0 \times 10^{-6} \text{ mol/s}$
- A2 Consider the following reaction:

$$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$$

 The rate of formation of NH_3 is 3.0 mL/min . The rate of consumption of H_2 is
 A. 1.5 mL/min B. 2.0 mL/min C. 4.5 mL/min D. 9.0 mL/min
- A2 Which of the following units can be used to express the rate of a chemical reaction?
 A. mL/g B. mol/L C. g/mol D. mol/min

Answers To Chemical Kinetics Study Guide

**Katherine J. Denniston, Denniston
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blends physical theory phenomenology and empiricism to provide a guide to the experimental practice and interpretation of reaction kinetics in solution It is suitable for courses in chemical kinetics at the graduate and advanced undergraduate levels This book will appeal to students in physical organic chemistry physical inorganic chemistry biophysical chemistry biochemistry pharmaceutical chemistry and water chemistry all fields concerned with the rates of chemical reactions in the solution phase

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Inorganic Chemistry J. E. House, 2012-10-30 This textbook provides essential information for students of inorganic chemistry or for chemists pursuing self study The presentation of topics is made with an effort to be clear and concise so that the book is portable and user friendly Inorganic Chemistry 2E is divided into five major themes structure condensed phases solution chemistry main group and coordination compounds with several chapters in each There is a logical progression from atomic structure to molecular structure to properties of substances based on molecular structures to behavior of solids etc The author emphasizes fundamental principles including molecular structure acid base chemistry coordination chemistry ligand field theory and solid state chemistry and presents topics in a clear concise manner There is a reinforcement of basic principles throughout the book For example the hard soft interaction principle is used to explain hydrogen bond strengths strengths of acids and bases stability of coordination compounds etc The book contains a balance of topics in theoretical and descriptive chemistry New to this Edition New and improved illustrations including symmetry and 3D molecular orbital representations Expanded coverage of spectroscopy instrumental techniques organometallic and bio inorganic chemistry More in text worked out examples to encourage active learning and to prepare students for their exams Concise coverage maximizes student understanding and minimizes the inclusion of details students are unlikely to use Discussion of elements begins with survey chapters focused on the main groups while later chapters cover the elements in greater detail Each chapter opens with narrative introductions and includes figures tables and end of chapter problem sets

American Chemical Journal Ira Remsen, Charles August Rouiller, 1895

Monthly Catalog of United States Government Publications, 1971

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