

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode $T \gg \theta_m$) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute $\frac{3}{2}R$ to the molar heat capacity. If we denote the number of active rotational modes by v_f^* (so for most molecules at normal temperatures $v_f^* = 2$ for linear molecules, and 3 for non-linear molecules), then

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The Journal of Physical Chemistry Wilder Dwight Bancroft,1917 Includes section New Books

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Journal of the Society of Chemical Industry Society of Chemical Industry (Great Britain),1912 Includes list of members 1882 1902 and proceedings of the annual meetings and various supplements

The Encyclopaedia Britannica ,1929

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The Encyclopedia Britannica James Louis Garvin,Franklin Henry Hooper,Warren E. Cox,1929

Chemical Engineering Catalog ,1918

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Bulletin of the Chemical Society of Japan Nihon Kagakkai, 1986 **Chemical News and Journal of Physical Science**, 1888 Physical Chemistry for Schools Henry John Horstman Fenton, 1916 **Industrial & Engineering Chemistry**, 1923 *The United States Catalog* Mary Burnham, Carol Hurd, 1928 *The Journal of Industrial and Engineering Chemistry*, 1923 Physical Chemistry Keith James Laidler, John H. Meiser, 1982 **Laboratory Manual of Organic Chemistry** Harry Linn Fisher, 1924 **The Lancet**, 1927

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